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ANALYTIC AND NUMERICAL CHEMISTRY ALGORITHMS FOR THE  
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David W. Goetz, et al

Air Force Weapons Laboratory  
Kirtland Air Force Base, New Mexico

June 1975

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# **ANALYTIC AND NUMERICAL CHEMISTRY ALGORITHMS FOR THE WORRY CODE**

**Worry Document No. 2**

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**June 1975**

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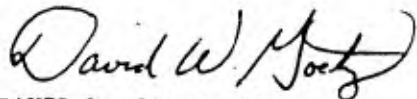
**AIR FORCE WEAPONS LABORATORY  
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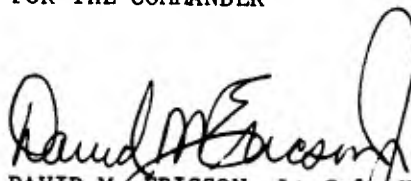


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20. ABSTRACT (Continue on reverse side if necessary and identify by block number) The WORRY Chemistry Package computes the nonequilibrium chemical behavior of a parcel of air in a quiescent or disturbed environment. Such a computation requires the solution of a set of nonlinear coupled differential equations derived from the set of reactions in which the atomic and molecular species of the parcel participate. The Chemistry Package may solve this set of differential equations by either of two methods: a standard numerical integration or an analytic algorithm. The procedures used in the analytic algorithm are described in detail and comparisons with the numerical solutions are shown.		

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SECTION I  
INTRODUCTION

Nuclear weapon effects codes such as the Air Force Weapons Laboratory (AFWL) WORRY optical-IR systems code require the repetitive and fast computation of non-equilibrium atmospheric chemical behavior at numerous points in a disturbed environment. Such computations necessitate the solution of a set of nonlinear coupled differential equations, often stiff, which describe the kinetic behavior of the chemical species of interest. The algorithm which is chosen for solution of these differential equations must be fast in computational time, reliable, and versatile in handling various initial conditions of species, temperature, and density.

At the inception of WORRY in late 1970-early 1971 a limited but adequate set of chemical species, table I, and their reactions, table II, were chosen to describe the terrestrial atmosphere from an optical/infrared viewpoint under both ambient and disturbed conditions. Several reliable algorithms were available for numerical integration of the differential equations defined by these species and their reactions. However, none were fast enough even in the most favorable situations and under adverse conditions e.g., at altitudes below 30 km, they often become inoperative. The answer to the need for a better algorithm was found by devising an analytic solution to the set of differential equations.

An analytic solution requires the making of simplifying assumptions that decouple the differential equations to a degree which leaves the remaining equations amenable to analytic solution. These assumptions obviously limit the generality and applicability of the method. The key to choosing a useful algorithm lies in defining a working range for the solutions and balancing uncertainties - in reactions, reaction rates, and solution assumptions. The analytic

algorithm for WORRY was so conceived. Since the chemical behavior to be described by the solution was limited by the structure of the WORRY code to points outside any fireballs, the two basic limiting assumptions of the analytic algorithms, (1) the molecular temperature does not exceed  $\sim 1200^{\circ}\text{K}$  and (2) the fractional ionization does not exceed  $\sim 0.1$  percent, were acceptable.

The continued development of the WORRY analytic algorithm has required a standard against which to judge its effectiveness. Therefore, a total chemistry package was developed for WORRY, containing both the analytic algorithm and a numerical algorithm to solve the same set of differential equations "exactly" and serve as a check on the "correctness" of the analytic solution. To ensure the comparability of outputs, both algorithms share the same inputs and the same subroutines which calculate initial conditions and reaction rates.

The analytic algorithm to be described in section IV is undergoing continual change. The description in section IV and the code listing (with external driver) found in section VII reflect the status of the chemistry package at the time of the writing of this document.

Table 1

WORRY SPECIES

H, Li, Be, Mg, Al, Si, Ti, Fe, Cu, U  
 OH, LiO, BeO, MgO, AlO, SiO, TiO, FeO, CuO, UO<sup>+</sup>  
 NO<sup>+</sup>, NO, NO<sub>2</sub>, N<sub>2</sub>O, CO, CO<sub>2</sub>, O<sub>3</sub>, H<sub>2</sub>O, CN, e<sup>-</sup>  
 N<sub>2</sub>, N<sub>2</sub>(A), N<sub>2</sub>(B), N(<sup>4</sup>S), N(<sup>2</sup>D), N(<sup>2</sup>P), O(<sup>3</sup>P), O(<sup>1</sup>D), O(<sup>1</sup>S), O<sup>-</sup>  
 N<sub>2</sub><sup>+</sup>, O<sub>2</sub><sup>+</sup>, N<sup>+</sup>, O<sup>+</sup>, N<sub>2</sub><sup>+</sup>(A), N<sub>2</sub><sup>+</sup>(B), O<sub>2</sub>, O<sub>2</sub>(<sup>1</sup>Δ), O<sub>2</sub>(<sup>1</sup>Σ), O<sub>2</sub>(<sup>3</sup>Σ)  
 H<sub>2</sub>, UO, NO(A), NO(B), NO(C), NO(D), CN(A)



TABLE II

## REACTIONS AND RATE CONSTANTS

$$k = a(T/300)^b \exp(-c/T) [M]^d$$

T = temperature (°K)

X = Li, Be, Mg, Al, Si, Ti, Fe, or Cu

	$\frac{a}{6.68 \times 10^{12}}$	$\frac{b}{0}$	$\frac{c}{3590}$	$\frac{d}{0}$
1. $N(^4S) + O_2 \rightarrow NO + O$ $N(^2P) + O_2 \rightarrow NO + O$				
2. $N(^2D) + O_2 \rightarrow NO + O$	$6.0 \times 10^{-12}$	.5	0	0
3. $N(^4S) + NO \rightarrow N_2 + O$ $N(^2P) + NO \rightarrow N_2 + O$	$5.1 \times 10^{-11}$	0	168	0
4. $N(^2D) + NO \rightarrow N_2 + O$	$6.0 \times 10^{-11}$	.5	0	0
5. $O + O_2 + M \rightarrow O_3 + M$	$2.8 \times 10^{-35}$	0	-900	1
6. $O + O_3 \rightarrow O_2 + O_2$	$2.0 \times 10^{-11}$	0	2400	0
7. $N_2^+ + e \rightarrow 0.99 N(^4S) +$ $0.99 N(^2D) + 0.02 N(^2P)$	$2.8 \times 10^{-7}$	-.2	0	0
8. $N_2^+ + O_2 \rightarrow O_2^+ + N_2$	$2.0 \times 10^{-10}$	0	0	0
9. $N_2^+ + O \rightarrow NO^+ + 0.5 N(^4S) +$ $0.5 N(^2D)$	$2.5 \times 10^{-10}$	0	0	0

Table II (Cont'd)

	a	b	c	d
10. $O_2^+ + e + 0.8 O(^3P) +$ $0.8 O(^1D) + 0.4 O(^1S)$	$2.2 \times 10^{-7}$	-0.7	0	0
11. $O^+ + O_2 \rightarrow O_2^+ + O(^3P)$	$2.0 \times 10^{-11}$	0	0	0
12. $NO^+ + e + 0.75 O(^3P) +$ $0.75 N(^2D) + 0.25 O(^1D) +$ $0.25 N(^4S)$	$5.0 \times 10^{-7}$	-1.2	0	0
13. $O_2 + N^+ + O_2^+ + 0.5 N(^4S) +$ $0.5 N(^2D)$	$2.5 \times 10^{-10}$	0	0	0
14. $O_2 + N^+ + NO^+ + O(^3P)$	$5.0 \times 10^{-10}$	0	0	0
15. $N_2 + O^+ \rightarrow NO^+ + N(^4S)$	$1.3 \times 10^{-12}$	0	0	0
16. $O^+ + NO \rightarrow NO^+ + O(^3P)$	$2.4 \times 10^{-11}$	0	0	0
17. $O_2^+ + NO \rightarrow NO^+ + O_2$	$8.0 \times 10^{-10}$	0	0	0
18. $N_2^+ + NO \rightarrow NO^+ + N_2$	$5.0 \times 10^{-10}$	0	0	0
19. $N^+ + NO \rightarrow NO^+ + 0.5 N(^4S) +$ $0.5 N(^2D)$	$8.0 \times 10^{-10}$	0	0	0

Table II (Cont'd)

	a	b	c	d
20. $O_2^+ + N \rightarrow NO^+ + O(^3P)$	$1.8 \times 10^{-10}$	0	0	0
21. $H + O_3 \rightarrow OH + O_2$	$2.6 \times 10^{-11}$	0	0	0
22. $OH + O \rightarrow H + O_2$	$5.0 \times 10^{-11}$	0	0	0
23. $O_2(^3\Sigma) + O_2 \rightarrow O_3 + O$	$1.0 \times 10^{-12}$	0	0	0
24. $O_3 + NO \rightarrow NO_2 + O_2$	$9.5 \times 10^{-13}$	0	1300	0
25. $O + NO + M \rightarrow NO_2 + M$	$1.3 \times 10^{-31}$	-2.5	0	1
26. $O + NO_2 \rightarrow NO + O_2$	$2.8 \times 10^{-11}$	0	550	0
27. $NO_2 + N \rightarrow N_2 + O_2$	$1.5 \times 10^{-11}$	0	0	0
28. $NO + N + M \rightarrow N_2O + M$	$1.0 \times 10^{-34}$	0	1000	1
29. $NO_2 + N \rightarrow N_2O + O$	$1.5 \times 10^{-11}$	0	0	0
30. $O_2(^1\Delta) + N_2O \rightarrow N_2 + O_3$	$1.0 \times 10^{-13}$	0	0	0
31. $N(^4S) + O_3 \rightarrow NO + O_2$	$3.4 \times 10^{-11}$	.5	1200	0
32. $N(^2P) \rightarrow N(^2D) + h\nu$	$7.78 \times 10^{-2}$	0	0	0
33. $O(^1S) \rightarrow O(^1D) + h\nu$	$6.99 \times 10^{-1}$	0	0	0
34. $O(^1D) \rightarrow O(^3P) + h\nu$	$6.25 \times 10^{-3}$	0	0	0

Table II (Cont'd)

	a	b	c	d
35. $N^+ + 0 \rightarrow 0^+ + N(^4S)$	$1.0 \times 10^{-12}$	0	0	0
36 - 49 for future expansion				
50 - 83 for metals reactions				
$X + O_2 \rightarrow XO + 0$				
$X + 0 + M \rightarrow XO + M$				
$X + O_3 \rightarrow XO + O_2$				
$XO + 0 \rightarrow X + O_2$				
$XO + O_3 \rightarrow X + O_2 + O_2$				
84 - 89 for future expansion				
90 $O_3 + h\nu \rightarrow O_2 + O(^3P)$				

SECTION II  
CHEMISTRY PACKAGE STRUCTURE

The WORRY Chemistry Package is structured as shown in the flow chart of figure 1. A requesting program calls DECAYR which either executes the analytic algorithm or transfers control to the numerical algorithm executive, DECAR, which employs subroutines FUNG and EVIL. The three routines and one entry point shared by both analytic and numerical solutions are ENTDEP, ENQDEP, RATE, and RRCRT. ENTDEP calculates the initial distribution of ions and neutrals produced by prompt energy deposition (ergs/cc), while the entry point ENQDEP partitions the continuous energy deposition (ion pairs/cc/sec) among the same set of ions and neutrals. For a given temperature and total density, the RATE subroutine fills the R array with the first-order, second-order, and pseudo-second-order reaction rate constants of the set of reactions in table II. RRCRT stores the species destruction rates and transforms the species formation rates as calculated by either algorithm into vibrational level formation rates to be used in the WORRY calculation of chemiluminescence. The remaining entries in figure 1 are functions used by DECAYR in its analytic solutions and will be elaborated in sections IV and V.

The chemistry package requires the following inputs for a problem:

1. The time,  $t_c$ , after instantaneous energy deposition for which the calculation is to be performed.
2. The initial, predeposition temperature ( $^{\circ}\text{K}$ ).
3. The instantaneous energy deposition (ergs/cm<sup>3</sup>) at  $t = 0$ .
4. The continuous energy deposition (ion pairs/cm<sup>2</sup>/sec) at  $t_c$ .
5. The time of calculation, as day or night.
6. Concentrations for any metals present.

7. An initial set of species concentrations at  $t = 0$  for ambient or previously insulted conditions. Ambient initialization may be accomplished by a prior call to the chemistry package with ambient inputs for 2, 3, and 4, a large value for  $t_c$ , and values for the concentrations of  $N_2$ ,  $O_2$ ,  $O$ ,  $O_3$ ,  $NO$ ,  $CO_2$ ,  $H$ , and  $OH$ .

The inputs as well as chemistry package outputs are accomplished through either the DECAIR calling arguments or COMMON blocks. The energy depositions (inputs), temperature (input/output), and time of calculation (input), are found in the DECAIR calling arguments. The variable DAY (input) is found in COMMON BURST and the DYMETL array (input) holds the metals concentrations (to be added to any preexisting concentrations). The CMD array (input/output) in COMMON XUX holds the concentrations of all species while the outputs of subroutine RRCRT are found in commoned arrays RR and RCRT.

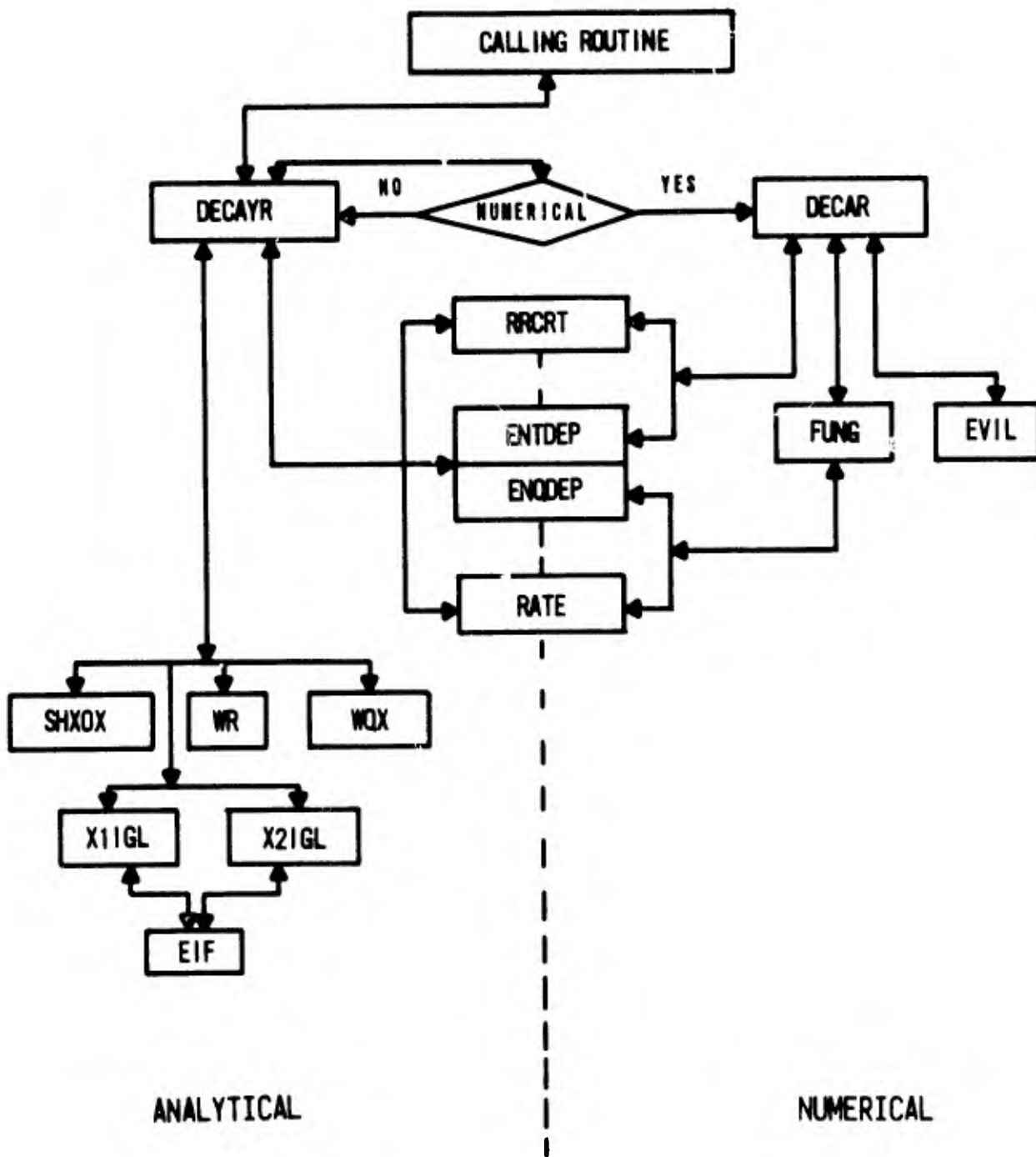


Figure 1. Flow Chart of Routines in WORRY Chemistry Package.

SECTION III  
NUMERICAL CHEMISTRY

In DECAR, the numerical chemistry algorithm begins by calling ENTDEP to initialize the species concentrations and choosing an initial time step, DT, of  $10^5/\text{CON}$ , where CON is the total number density. The numerical integration is performed by a modified Kenneshea technique. The differential equation for each species is assumed to have the form  $d[x]/dt = P - R[x]$ , where  $[x]$  is the species concentration, P is the production rate in  $\text{cc}^{-1} \text{sec}^{-1}$ , and R is the removal rate in  $\text{sec}^{-1}$ . P and R are assumed to be independent of  $[x]$ ; and if they are assumed to be constant over the time interval DT, the solution may be written.

$$[x(T+DT)] = ([x(T)] - P/R)E^{-RT} + P/R \quad (1)$$

P and R are evaluated at time T;  $[x]$  is then evaluated at (T+DT) and used to evaluate P and R at (T+DT). These values of P and R are then used with  $[x(T)]$  to re-evaluate  $[x]$  at (T+DT). It is required that  $|[x(T+DT)] - [x(T)]|/[x(T)]$  be less than  $\epsilon$  for each species, where  $\epsilon = 10^{-4}$  at the present time. The above process of requiring convergence is repeated three times, and if the requirement is not met, DT is halved. If the requirement is met, DT is doubled for the next step. This process is continued until the desired calculational time, END, is reached. When END is reached, RRCRT is called and the algorithm is complete.

The subroutine FUNG is used to evaluate P and R, and the subroutine EVIL is used to calculate  $[x(T+DT)]$  from  $[x(T)]$ , P, and R. In addition to the species



concentrations, the internal kinetic energy of the gas is carried as a variable in the CMD array. The release and absorption of energy for each reaction is calculated in FUNG, and the energy is updated along with the other species. The updated energy is used to calculate a new temperature for each time step and a call to RATE is made to recalculate the reaction rate constants if necessary.

SECTION IV  
ANALYTIC CHEMISTRY

As in the numerical chemistry, the analytic algorithm found in subroutine DECAYR begins by calling ENTDEP to update the CMD species array for the prompt energy deposition production of ions and neutrals at  $t = 0$ . The resultant updated temperature from ENTDEP is then used in a call to the RATE routine which fills the R array with all necessary rates constants for the reactions of table 2. As a final initialization step, the algorithm calculates the total energy of the sample volume using a table of heats of formation, the CMD array of species concentrations, and the total internal kinetic energy of the volume as calculated in ENTDEP.

The first assumption in the analytic algorithm is the separation of the fast-charged species from the slower neutral species chemistry. This separation and the following treatment are acceptable under (and lead to) the two basic assumptions of the analytic algorithm listed in section I. The faster charged species reactions are treated first and the resultant heating is calculated before the neutral species chemistry is begun. The charged species--  $N^+$ ,  $O^+$ ,  $N_2^+$ ,  $O_2^+$ , and  $NO^+$  --react primarily through charge exchange and recombination reactions as found in reactions 7-20 and 35. The following description tracks the computer coding closely through the charged species and neutral species chemistries.

1. NITROGEN ION ( $N^+$ )

$N^+$  is at the top of the charge exchange chain and has no production from other species reactions. Assuming, as will be done successfully in many cases throughout the analytic treatment, that the concentrations of the species with

which  $N^+$  reacts are constant, the  $N^+$  loss mechanisms of reactions 13, 14, 19 and 35 result in the homogeneous differential equation for  $[N^+]$ .

$$d[N^+]/dt = - \left\{ k_{13}[O_2] + k_{14}[O_2] + k_{19}[NO] + k_{35}[O] \right\} [N^+] \quad (2)$$

The sum of constant quantities multiplying  $[N^+]$  is defined as  $\beta_N$  and the equation is rewritten as

$$d[N^+]/dt = - \beta_N [N^+] \quad (3)$$

to which the solution is a simple exponential decay.

$$[N^+] = [N^+]_0 e^{-\beta_N t} \quad (4)$$

The full differential equation for  $[N^+]$  including the continuous source is

$$d[N^+]/dt = Q_N + - \beta_N [N^+] \quad (5)$$

where  $Q_N$  is the production rate of  $[N^+]$  from the continuous energy deposition. The particular solution to this inhomogeneous equation form is calculated in the Chemistry Package by FUNCTION WQX, which is described in Section V.

The final concentration of  $N^+$  at the calculational time is taken as the sum of the solution,  $[N^+]_t$ , of Eq (4) and the particular solution for Eq (5) where  $[N^+]_0$  in the later case is assumed to be zero. This procedure is general throughout the analytic treatment. Each species is treated twice - once without and once with the continuous energy deposition considered- and the results are either added or the maximum is

taken depending upon the species and the specific mathematical treatment and boundary conditions assumed.

## 2. OXYGEN ION ( $O^+$ )

$O^+$  is handled in a manner similar to  $N^+$  except that the differential equation for  $[O^+]$  includes a production term from the charge exchange of  $N^+$  with  $O$ . Consideration of Reactions 11, 15, 16, and 35 leads to the inhomogeneous equation for the instantaneous solution.

$$d[O^+]/dt = - \{k_{11}[O_2] + k_{15}[N_2] + k_{16}[NO]\} [O^+] + k_{35}[N^+][O] \quad (6)$$

The equation is rewritten by defining the constants  $\beta_0$  and  $\gamma$ .

$$d[O^+]/dt = - \beta_0 [O^+] + \gamma [N^+] \quad (7)$$

Substituting Eq (4) for  $[N^+]$  gives

$$d[O^+]/dt = - \beta_0 [O^+] + \gamma [N^+]_0 e^{-\beta_N t} \quad (8)$$

to which the general solution is of the form

$$[O^+] = \alpha e^{-\beta_0 t} + [\gamma [N^+]_0 / (\beta_0 - \beta_N)] e^{-\beta_N t} \quad (9)$$

The value of  $\alpha$  is determined from initial conditions and the solution takes the form

$$[O^+] = \{[O^+]_0 - \gamma [N^+]_0 / (\beta_0 - \beta_N)\} e^{-\beta_0 t} + \{\gamma [N^+]_0 / (\beta_0 - \beta_N)\} e^{-\beta_N t} \quad (10)$$

If  $y = \gamma [N^+]_0 / (\beta_0 - \beta_N)$  the final form of the instantaneous solution is

$$[O^+] = ([O^+]_0 - y) e^{-\beta_0 t} + y e^{-\beta_N t} \quad (11)$$

The solution for  $[O^+]$  including the continuous source is found as for  $[N^+]$  from the equation

$$d[O^+]/dt = Q - \beta_0[O^+] \quad (12)$$

where  $Q = Q_{O^+} + \gamma[N^+]_t$  and  $[O^+]_0$  is assumed to be zero.

The final concentration of  $O^+$  at the calculational time is the sum of the instantaneous and continuous energy deposition solutions.

### 3. ELECTRON CONCENTRATION

Before proceeding with the positive molecular ions, an intermediate update is made to the electron concentration. The same calculational scheme is repeated after the positive molecular ions have been treated to give the final electron concentration. (At present,  $n_e$  is considered as the total negative charge concentration in an equal  $\alpha$  approximation.)

Consideration of the continuous source production rate is postponed as usual, and Reactions 7, 10, and 12 define the electron concentration behavior

$$dn_e/dt = - \left\{ k_7[N_2^+] + k_{10}[O_2^+] + k_{12}[NO^+] \right\} n_e \quad (13)$$

where the approximation has been made that atomic ions do not recombine—a valid approximation due to their very slow recombination rates compared to the molecular ions. An average recombination rate,  $\alpha_r$ , is defined

$$\alpha_r = \frac{k_7[N_2^+] + k_{10}[O_2^+] + k_{12}[NO^+]}{[N_2^+] + [O_2^+] + [NO^+] + [O^+] + [N^+]}$$

and Eq (13) is replaced by

$$dn_e/dt = - \alpha_r n_e P \quad (14)$$

where  $P$  is the number of molecular positive ions,

$P = n_e - ([O^+] + [N^+])$ . Defining  $[A^+] = [O^+] + [N^+]$ , Eq (14) is rewritten as

$$dn_e/dt = -\alpha_r n_e (n_e - [A^+]) \quad (15)$$

and if  $[A^+]$  is treated as a constant the solution for  $n_e$  is

$$n_e = [A^+] + (n_o - [A^+]) / (1 + \alpha_r' n_o t) \quad (16)$$

where  $\alpha_r' = \alpha_r [\exp(\alpha_r [A^+] t) - 1] / (\alpha_r [A^+] t)$  and  $n_o$  is the initial electron concentration at  $t = 0$ . This expression for  $n_e$  is exact in the limits of  $[A^+]$  equal to zero and infinity, and from comparisons with numerical solutions, Eq (16) appears to be valid over a wide range of  $[A^+]$  values.

The continuous source production of negative charge is handled simply by the solution of the equation

$$dn_e/dt = Q - \alpha_r n_e^2 \quad (17)$$

where the ion-pair production rate,  $Q$ , has been added to Eq (14) and the approximation,  $P = n_e$ , has been made.

The solution to Eq (17) is

$$n_e = \sqrt{\frac{Q}{\alpha_r}} \frac{(1 + a e^{-2\sqrt{\alpha_r Q} t})}{(1 - a e^{-2\sqrt{\alpha_r Q} t})} \quad (18)$$

where  $a = (n_o - \sqrt{\frac{Q}{\alpha_r}}) / (n_o + \sqrt{\frac{Q}{\alpha_r}})$ . This solution is contained in the Chemistry Package as FUNCTION WR.

The final electron concentration  $n_e$  at the calculational time is the maximum of the instantaneous and continuous source solutions.

#### 4. MOLECULAR NITROGEN ION ( $N_2^+$ )

$N_2^+$  is considered next in the charge exchange chain of positive ions. The differential equation for  $[N_2^+]$  is derived from Reactions 7, 8, 9, and 18

$$d[N_2^+]/dt = -k_7 n_e [N_2^+] - \{k_8[O_2] + k_9[O] + k_{18}[NO]\} [N_2^+] \quad (19)$$

Equation (19) may be rewritten by defining  $\beta_{N_2} = k_8[O_2] + k_9[O] + k_{18}[NO]$  and substituting Eq (16) for  $n_e$

$$d[N_2^+]/dt = - \left\{ k_7[A^+] + k_7(n_o - [A^+]) / (1 + \alpha'_r n_o t) + \beta_{N_2} \right\} [N_2^+] \quad (20)$$

If  $\alpha'_r$  and  $\beta_{N_2}$  are treated as constants and Eq (4) and Eq (11) define

$[A^+] = [O^+] + [N^+]$ , the instantaneous solution for  $N_2^+$  is

$$\begin{aligned} [N_2^+] = & [N_2^+]_o \exp \left\{ -\beta_{N_2} t - k_7 \left( ([O^+]_o - y) / \beta_o \right) (1 - e^{-\beta_o t}) \right. \\ & - k_7 \left( ([N^+]_o + y) / \beta_N \right) (1 - e^{-\beta_N t}) + (k_7 / (\alpha'_r n_o)) \left( ([O^+]_o - y) e^{A_1} (E_1(A_1) - E_1(B_1)) \right. \\ & \left. \left. + ([N^+]_o + y) e^{A_2} (E_1(A_2) - E_1(B_2)) \right) \right\} / (1 + \alpha'_r n_o t)^{k_7 / \alpha'_r} \end{aligned} \quad (21)$$

where  $A_1 = \beta_o / (\alpha'_r n_o)$ ,  $B_1 = A_1 (1 + \alpha'_r n_o t)$ ,  $A_2 = \beta_N / (\alpha'_r n_o)$ ,  $B_2 = A_2 (1 + \alpha'_r n_o t)$ , and

$E_1(A)$  is the exponential integral  $\int_A^\infty \frac{e^{-x}}{x} dx$  which is calculated by

FUNCTION EIF. Equation (21) has been modified during the development of the analytic treatment to assure the best possible agreement with numerical solutions. In particular, the denominator  $(1 + \alpha'_r n_o t)^{k_7 / \alpha'_r}$  has been replaced by  $(1 + \alpha'_r n_o t)$ , resulting in a more even performance of the solution over a wide range of  $\alpha'_r$  values.

The inclusion of a continuous source term in Eq (19) results in the full equation

$$d[N_2^+]/dt = Q_{N_2^+} - k_7 n_e [N_2^+] - \beta_{N_2^+} [N_2^+] \quad (22)$$

which may be regrouped if  $n_e$  is assumed to be a constant

$$d[N_2^+]/dt = Q_{N_2^+} - \beta [N_2^+] \quad (23)$$

where  $\beta = k_7 n_e + \beta_{N_2^+}$ .

The solution to Eq (23) is found as for  $N^+$  and  $O^+$  using the FUNCTION WQX with  $[N_2^+]_0 = 0$ .

The final concentration of  $N_2^+$  at  $t_c$  is considered to be the maximum of the instantaneous and continuous source solutions.

#### 5. MOLECULAR OXYGEN ION ( $O_2^+$ )

Next in the charge exchange chain,  $O_2^+$  is involved in Reactions 8, 10, 11, 13, 17, and 20. Of these, Reactions 8, 11 and 13 produce  $[O_2^+]$  and are treated as increasing  $[O_2^+]$  at  $t = 0$  by an amount

$$x_{O_2^+} = k_8 [O_2] \int_0^t [N_2^+] dt' + k_{11} [O_2] \int_0^t [O^+] dt' + k_{13} [O_2] \int_0^t [N^+] dt' \quad (24)$$

where  $I_{O^+} = \int_0^t [O^+] dt'$  and  $I_{N^+} = \int_0^t [N^+] dt'$  can be found by integrating

Eq (11) and Eq (4) respectively.



The integral  $I_{N_2^+} = \int_0^t [N_2^+] dt'$  can not be performed exactly, therefore, an integrable form is assumed

$$[N_2^+] = [N_2^+]_0 e^{-\beta \frac{N_2}{e} t} / (1 + \alpha_r n_o t) \quad (25)$$

where  $\beta \frac{N_2}{e}$  is calculated by setting Eq (25) equal to the results of Eq (21). The integral may now be evaluated

$$I_{N_2^+} = [N_2^+]_0 (e^{A_3 / (\alpha_r n_o)}) [E_1(A_3) - E_1(B_3)] \quad (26)$$

where  $A_3 = \beta \frac{N_2}{e} / (\alpha_r n_o)$  and  $B_3 = A_3 (1 + \alpha_r n_o t)$  and Eq (24) is condensed to

$$x_{O_2^+} = [O_2^+] \left\{ k_8 I_{N_2^+} + k_{11} I_{O^+} + k_{13} I_{N^+} \right\} \quad (27)$$

The remaining Reactions 10, 17, and 20 define the differential equation for  $[O_2^+]$

$$d[O_2^+] / dt = -k_{10} [O_2^+] n_e - k_{17} [NO] [O_2^+] - k_{20} [N] [O_2^+] \quad (28)$$

or substituting Eq (16) and defining  $\beta_{O_2} = k_{17} [NO] + k_{20} [N]$

$$d[O_2^+] / dt = - \left\{ k_{10} \left[ [A^+] + (n_o - [A^+]) / (1 + \alpha_r' n_o t) \right] + \beta_{O_2} \right\} [O_2^+] \quad (29)$$

Treating Eq (29) in the same manner as Eq (20), the final instantaneous solution for  $[O_2^+]$  is

$$[O_2^+] = [O_2^+]_0 \exp \left\{ -\beta_{O_2} t - k_{10} I_{A^+} + k_{10} / (\alpha_r' n_o) \left( ([O_2^+]_0 - y) e^{A_1} (E_1(A_1) - E_1(B_1)) \right) \right\}$$

$$+ ([N^+]_0 + y) e^{A_2(E_1(A_2) - E_1(B_2))} \left\{ / (1 + \alpha'_r n_o t)^{k_{10}/\alpha'_r} \right. \quad (30)$$

where  $I_{A^+} = I_{O^+} + I_{N^+} = (([O^+]_0 - y)/\beta_O)(1 - e^{-\beta_O t}) + (([N^+]_0 + y)/\beta_N)(1 - e^{-\beta_N t})$ ;

$A_1$ ,  $B_1$ ,  $A_2$ , and  $B_2$  have been defined; and  $[O_2^+]_0$  includes  $X_{O_2^+}$ .

The above treatment has been modified to take into account some of the errors introduced by the approximations made. Since  $[O_2^+]_0$  is increased by an amount  $X_{O_2^+}$  at  $t = 0$  allowing that additional concentration to recombine from  $t = 0$  and not from the actual time of creation, Eq (27) is modified to correct for the recombination of  $N_2^+$  by the factor

$$C_{N_2} = [N_2^+]_0 / \left\{ [N_2^+] + (k_8[O_2] + k_9[O] + k_{18}[NO])I_{N_2^+} \right\} \quad (31)$$

where the latter term is the  $N_2^+$  lost to all processes except recombination.

Equation (27) then has the form

$$X_{O_2^+} = [O_2] \left\{ k_8 C_{N_2} + I_{N_2^+} + k_{11} I_{O^+} + k_{13} I_{N^+} \right\} \quad (32)$$

Equation (30) has also been altered to improve agreement with numerically calculated solutions. If Eq (30) is rewritten as

$$[O_2^+] = [O_2^+]_0 \exp \left\{ -\beta_{O_2} t - k_{10}(I_{A^+} - I) \right\} / (1 + \alpha'_r n_o t)^{k_{10}/\alpha'_r} \quad (33)$$

the modification to Eq (33) is the denominator transformation

$$(1 + \alpha'_r n_o t)^{k_{10}/\alpha'_r} \rightarrow (1 + \alpha'_r n_o t (\exp(-k_{10}(I_{A^+} - I))))^{k_{10}/\alpha'_r}$$

The continuous source calculation for  $O_2^+$  is handled in a similar manner to  $N_2^+$ . The full differential equation for  $O_2^+$

$$d[O_2^+]/dt = Q_{O_2^+} + \left\{ k_8[N_2^+] + k_{11}[O^+] + k_{13}[N^+] \right\} [O_2] - \left\{ k_{10}n_e + \beta_{O_2} \right\} [O_2^+] \quad (34)$$

can be regrouped into the familiar form

$$d[O_2^+]/dt = Q - \beta [O_2^+] \quad (35)$$

where  $Q = Q_{O_2^+} + \left\{ k_8[N_2^+] + k_{11}[O^+] + k_{13}[N^+] \right\} [O_2]$

and  $\beta = k_{10}n_e + \beta_{O_2}$ . Eq (35) is solved in the usual way using FUNCTION WQX with  $[O_2^+]_0 = 0$ .

The final concentration of  $O_2^+$  at  $t_c$  is considered to be the maximum of the instantaneous and continuous source solutions.

#### 6. NITRIC OXIDE ION ( $NO^+$ )

The final species in the charge exchange chain,  $NO^+$ , is involved in Reactions 9, 12, 14, 15, 16, 17, 18, 19, and 20. All these reactions except 12 may be viewed as increasing  $[NO^+]$  at  $t = 0$  by an amount

$$X_{NO^+} = k_9[O]I_{N_2^+} + k_{14}[O_2]I_{N^+} + k_{15}[N_2]I_{O^+} + k_{16}[NO]I_{O^+} + k_{17}[NO]I_{O_2^+} + k_{18}[NO]I_{N_2^+} + k_{19}[NO]I_{N^+} + k_{20}[N]I_{O_2^+} \quad (36)$$

where  $I_{O_2^+} = \int_0^t [O_2^+] dt'$ . In analogy to the calculation of  $I_{N_2^+}$ ,

an integrable form is assumed for  $[O_2^+]$

$$[O_2^+] = [O_2^+]_0 e^{-\beta_{O_2} t} / (1 + \alpha_r n_e t) \quad (37)$$

where  $\beta_{O_2}$  is calculated by setting Eq (37) equal to the results of Eq (30)

Thus

$$I_{O_2^+} = [O_2^+]_0 e^{A_4 / (\alpha_r n_o)} (E_1(A_4) - E_1(B_4))$$

where  $A_4 = \beta \frac{O_2}{e} / (\alpha_r n_o)$  and  $B_4 = A_4(1 + \alpha_r n_o t)$ . (38)

This approach leaves the differential equation for the instantaneous solutions of  $[NO^+]$  in the form

$$d[NO^+] / dt = - k_{12} n_e [NO^+] \quad (39)$$

Substituting Eq (16) for  $n_e$  and assuming  $\alpha_r'$  a constant, the equation is integrated to give the solution

$$[NO^+] = [NO^+]_0 \exp\{-k_{12} I_{A^+} + k_{12} / (\alpha_r' n_o) [([O^+]_0 - y) e^{A_1} (E_1(A_1) - E_1(B_1)) + ([N^+]_0 + y) e^{A_2} (E_1(A_2) - E_1(B_2))]\} / (1 + \alpha_r' n_o t)^{k_{12} / \alpha_r'} \quad (40)$$

where  $A_1, B_1, A_2,$  and  $B_2$  have been defined previously, and  $[NO^+]_0$  includes  $X_{NO^+}$ .

Again, as for  $[O_2^+]$ , Eq (36) and Eq (40) must be modified to correct approximation errors. In Eq (36), all terms involving  $I_{N_2^+}$  are multiplied by the factor  $C_{N_2}$  and all involving  $I_{O_2^+}$ , by the factor

$$C_{O_2} = [O_2^+]_0 / ([O_2^+] + (k_{17}[NO] + k_{20}[N])I_{O_2^+}) \quad (41)$$

to correct for the fact that these concentrations did not recombine from  $t = 0$ .

Eq (40) is not only modified but is also passed through an iteration process along with the continuous source solution. First, if Eq (40) is rewritten as

$$[\text{NO}^+] = [\text{NO}^+]_0 \exp \left\{ -k_{12} (I_{A^+} - I) \right\} / (1 + \alpha'_r n_o t)^{k_{12}/\alpha'_r} \quad (42)$$

the denominator is modified to give the first iteration instantaneous solution form

$$[\text{NO}^+] = [\text{NO}^+]_0 \exp \left\{ -k_{12} (I_{A^+} - I) \right\} / (1 + \alpha'_r n_o t (\exp(-k_{12} (I_{A^+} - I)))) \quad (43)$$

The first iteration continuous source solution solves the full differential equation for  $[\text{NO}^+]$  which is analogous to Eq (34) and may be reduced to

$$d[\text{NO}^+]/dt = Q - \beta[\text{NO}^+] \quad (44)$$

which is solved using FUNCTION WQX with  $[\text{NO}^+]_0 = 0$  and  $\beta = k_{12} n_e$ .

The maximum of the instantaneous and continuous source solutions of Eq (43) and Eq (44) constitute the first iteration solution of  $[\text{NO}^+]$ .

Having calculated concentrations at  $t_c$  for all the positive molecular ions, a new  $\alpha_r$  and  $\alpha'_r$  are calculated and the formalism for the calculation of  $n_e$  in Section IV-3 is repeated to update the electron (i.e. negative ion) concentration at  $t_c$ .

The second iteration on  $[\text{NO}^+]$  then uses the new  $\alpha'_r$  to recalculate the instantaneous solution of Eq (42) modified as

$$[\text{NO}^+] = [\text{NO}^+]_0 \exp \left\{ -k_{12} (I_{A^+} + I) \right\} / (1 + \alpha'_r n_o t (\exp(-k_{12} (I_{A^+} - I))))^{k_{12}/\alpha'_r} \quad (45)$$

with  $I$  not being recalculated. Likewise, the continuous solution is recalculated using the new  $n_e$ . Finally, the maximum of the two solutions is taken to give the ultimate  $[\text{NO}^+]$  concentration at  $t_c$ .

7. SPECIES AND TEMPERATURE UPDATING

Throughout the above treatment all neutral species concentrations were assumed to be constant. This approximation alone is insufficient, however, and intermediate updating of some neutral species concentrations is accomplished between the solutions for  $[N_2^+]$  and  $[O_2^+]$  (i.e., Sections IV-4 and IV-5). Likewise, the neutral species concentrations and the temperature are ultimately modified after the entire treatment to account for production by the fast ion chemistry.

a. Intermediate Updating of NO, N, and O

Intermediate updating is accomplished for  $[N]$ ,  $[O]$ , and  $[NO]$ .  $[N]$  and  $[O]$  created or destroyed by an ion-neutral reaction are accounted for by additions to their initial concentrations,  $[N]_0$  or  $[O]_0$

$$[Y]_0 = [Y]_0 \pm bk[X]I_{m^+} \tag{46}$$

where  $I_{m^+}$  is either  $I_{N^+}$ ,  $I_{O^+}$ , or  $I_{N_2^+}$ ,  $b$  is a branching ratio,  $k$  is the reaction rate constant, and  $X$  is the reacting neutral whose concentration is assumed to be constant. The increase in  $[N]$  due to the dissociative recombination of  $N_2^+$  additionally requires the more complex integral

$k_7 \int_0^t n_e [N_2^+] dt'$  where neither term of the product can be assumed to be constant. This integral is calculated by substituting Eq (16) and Eq (25) for  $n_e$  and  $[N_2^+]$  respectively and solving.

$$I_{N_2^+}^{n_e} = \int_0^t [N_2^+] n_e dt' = \int_0^t \frac{[N_2^+]_0 e^{-\beta e^{N_2} t'}}{1 + \alpha_r n_0 t'} \left\{ [A^+] + \frac{n_0 - [A^+]}{1 + \alpha_r' n_0 t'} \right\} dt' \tag{47}$$

$$I_{N_2^+}^{n_e} = [N_2^+]_0 \left\{ \int_0^t \frac{[A^+] e^{-\beta e^{N_2} t'}}{1 + \alpha_r n_0 t'} dt' + \int_0^t \frac{(n_0 - [A^+]) e^{-\beta e^{N_2} t'}}{(1 + \alpha_r n_0 t')(1 + \alpha_r' n_0 t')} dt' \right\} \tag{48}$$

The denominator in the second integral is approximated by  $(1 + \gamma n_0 t')^2$

where the constant  $\gamma = \sqrt{\alpha_r \alpha_r'}$ . Integrating yields

$$\begin{aligned}
 I_{N_2}^{n_e+} = & [N_2^+]_0 \left\{ \frac{([O^+]_0 - y)}{\alpha_r n_0} e^{A_5} (E_1(A_5) - E_1(B_5)) + \frac{([N^+]_0 + y)}{\alpha_r n_0} e^{A_6} (E_1(A_6) - E_1(B_6)) \right. \\
 & + \frac{1}{\gamma} \left[ \left(1 - \frac{e^{-\beta_e N_2 t}}{1 + \gamma n_0 t}\right) - A_7 e^{A_7} (E_1(A_7) - E_1(B_7)) \right] - \frac{([O^+]_0 - y)}{\gamma n_0} \left[ \left(1 - \frac{e^{-\delta t}}{1 + \gamma n_0 t}\right) \right. \\
 & \left. \left. - A_8 e^{A_8} (E_1(A_8) - E_1(B_8)) \right] - \frac{([N^+]_0 + y)}{\gamma n_0} \left[ \left(1 - \frac{e^{-\sigma t}}{1 + \gamma n_0 t}\right) - A_9 e^{A_9} (E_1(A_9) - E_1(B_9)) \right] \right\} \quad (49)
 \end{aligned}$$

where  $\delta = \beta_e N_2 + \beta_o$ ,  $\sigma = \beta_e N_2 + \beta_N$ ,  $A_5 = \frac{\delta}{\alpha_r n_0}$ ,  $B_5 = A_5(1 + \alpha_r n_0 t)$ ,  $A_6 = \frac{\sigma}{\alpha_r n_0}$

$B_6 = A_6(1 + \alpha_r n_0 t)$ ,  $A_7 = \frac{\beta_e N_2}{\gamma n_0}$ ,  $B_7 = A_7(1 + \gamma n_0 t)$ ,  $A_8 = \frac{\delta}{\gamma n_0}$ ,  $B_8 = A_8(1 + \gamma n_0 t)$ ,

$A_9 = \frac{\sigma}{\gamma n_0}$ , and  $B_9 = A_9(1 + \gamma n_0 t)$ .

Having calculated  $I_{N_2}^{n_e}$ ,  $[N]$  can be updated using the recombination

term  $k_7 I_{N_2}^{n_e+}$ . However, since several approximations were made in the calculation, a further modification is necessary. The total amount of  $N_2^+$  that has been removed by the several reactions in which it participates is

$$\Delta N_2^+ = [N_2^+]_0 - [N_2^+]_t \quad (50)$$

which if no approximations had been made in calculating  $I_{N_2}^{n_e+}$  and  $I_{N_2}^{n_e}$  would be equivalent to

$$A_{N_2} = (k_8[O_2] + k_9[O] + k_{18}[NO]) I_{N_2}^{n_e+} + k_7 I_{N_2}^{n_e} \quad (51)$$

Since the two terms will not be equal, the concentration of  $N_2^+$  which has recombined is calculated as the total change in  $[N_2^+]$  multiplied by the ratio of that which has recombined to the total that has reacted. Thus the concentration of  $N_2^+$  which has recombined is

$$R_{N_2^+} = \Delta N_2^+ \left( k_7 \frac{I_{N_2^+}^n}{A_{N_2^+}} \right) \quad (52)$$

which is used along with the appropriate branching ratios to update the N concentrations.

The intermediate updating of [NO] is accomplished in a very approximate fashion. Reactions 1, 2, 3, 4, 24, 25, 26, 28, and 31 involve [NO], but only the first four are considered, resulting in the differential equation

$$\frac{d[NO]}{dt} = [NSP](k_1[O_2] - k_3[NO]) + [N(^2D)](k_2[O_2] - k_4[NO]) \quad (53)$$

where  $[NSP] = [N(^4S)] + [N(^2P)]$ . The number of variables in Eq (53) can be reduced by assuming that an increase in [NO] is reflected as a decrease in  $[O_2]$  such that

$$[O_2] = [O_2]_0 - ([NO] - [NO]_0) \quad (54)$$

Substituting into Eq (53)

$$\frac{d[NO]}{dt} = [NSP] \left\{ k_1([O_2]_0 + [NO]_0) - (k_1+k_3)[NO] \right\} + [N(^2D)] \left\{ k_2([O_2]_0 + [NO]_0) - (k_2+k_4)[NO] \right\} \quad (55)$$



If reactions involving  $N(^2D)$  are more rapid than those of  $N(^4S)$  and  $N(^2P)$ , the two halves of Eq (55) may be solved sequentially.

First, the solution for  $[N(^2D)]$  is assumed to have the form

$$[N(^2D)] = [N(^2D)]_0 e^{-\beta_1 t} \quad (56)$$

where  $\beta_1 = k_2[O_2]_0 + k_4[NO]_0$ . And the solution to the second half of Eq (55) is

$$[NO]_1 = k_2([O_2]_0 + [NO]_0)/(k_2+k_4) - \left\{ k_2([O_2]_0 + [NO]_0)/(k_2+k_4) - [NO]_0 \right\} \exp \left\{ -(k_2+k_4)[N(^2D)]_0 (1-e^{-\beta_1 t})/k_1 \right\} \quad (57)$$

The result of Eq (57) is then used as the initial condition for  $[NO]$  in the solution to the first half of Eq (55). Thus, the intermediate updated concentration for NO is

$$[NO] = k_1([O_2]_0 + [NO]_0)/(k_1+k_3) - \left\{ k_1([O_2]_0 + [NO]_0)/(k_1+k_3) - [NO]_1 \right\} \exp \left\{ -(k_1+k_3)[NSP]_0 (1-e^{-\beta_2 t})/\beta_2 \right\} \quad (58)$$

where  $\beta_2 = k_1[O_2]_0 + k_3[NO]_1$  and the time dependence of  $[NSP]$  has the form of Eq (56).

#### b. FINAL UPDATING OF N AND O

After completion of the fast ion chemistry, the concentrations of N and O are updated to account for production by the reactions of  $O_2^+$  and  $NO^+$ . Procedures similar to those for the intermediate updating are followed. In particular for the recombination of  $O_2^+$ , the integral is needed

$$I_{O_2^+}^{n_e} = \int_0^t [O_2^+] n_e dt' = [O_2^+]_0 \left\{ \int_0^t \frac{[A^+] e^{-\beta_e^{O_2^+} t'}}{(1 + \alpha_{r_{O_2^+}} t')} dt' + \int_0^t \frac{(n_0 - [A^+]) e^{-\beta_e^{O_2^+} t'}}{(1 + \gamma_{r_{O_2^+}} t')^2} dt' \right\} \quad (50)$$

which can be solved in analogy to  $N_2^+$  in Eq (48) to give the solution of Eq (49) where  $[O_2^+]_0$  is substituted for  $[N_2^+]_0$  and  $\beta_e^{O_2^+}$  for  $\beta_e^{N_2^+}$ . A similar integral for  $[NO^+]$ , however is not required since the only loss mechanism for  $NO^+$  is recombination and the integral is simply

$$k_{12} \int_0^t n_e [NO^+] dt' = \Delta NO^+ = [NO^+]_0 - [NO^+]_t \quad (60)$$

Again in analogy to the  $N_2^+$ , the recombination integral for  $O_2^+$  is corrected for the approximations made throughout the calculation and the final concentration of  $O_2^+$  which has recombined is calculated as

$$R_{O_2^+} = \Delta O_2^+ (k_{10} I_{O_2^+}^{n_e} / A_{O_2^+}) ([O_2^+]_0' / [O_2^+]_0) \quad (61)$$

where a second correction term ( $[O_2^+]_0' / [O_2^+]_0$ ) has been added to account for the effects of  $[O_2^+]_0$  having been increased by  $X_{O_2^+}$ , Eq (32), and the approximations made there. Similarly, the final concentration of  $NO^+$  which has recombined is calculated as

$$R_{NO^+} = \Delta NO^+ ([NO^+]_0' / [NO^+]_0) \quad (62)$$

where the correction term has been added to account for the approximations made in updating  $[NO^+]_0$  with  $X_{NO^+}$ .

The branching ratios of Reactions 10 and 12 are used with  $R_{O_2^+}$  and  $R_{NO^+}$  respectively, to update the N and O concentrations.

## c. FINAL UPDATING OF TEMPERATURE AND REACTION RATES

The final step in the fast ion chemistry treatment is the recomputation of the gas temperature and reaction rates. For the temperature calculation, the specific heat of the sample volume is recalculated as is the total heat of formation for all species in the volume. The internal energy of the volume is then recalculated as its former value plus the change in the total heat of formation for all species. This new internal energy is finally converted to an updated gas temperature using the new specific heat. The new temperature is then used in RATE to recalculate the reaction rates for all reactions of Table II.

## 8. NEUTRAL SPECIES CHEMISTRY

The slower neutral chemistry is much simpler in its mathematics, but is intricate in the ordering of species and reactions for solution. The subsequent discussion, therefore, follows the programmed logic rather than segmentation by species.

In general, the treatment follows four steps. In the first, the continuous source production rates are neglected and  $[O]$ ,  $[O_2(^3\Sigma)]$ ,  $[N]$ ,  $[NO]$ , and  $[O_3]$  are addressed. In the second the continuous source rates are included in the solutions for most species. The reactions involving  $[OH]$  and  $[H]$  are treated separately in the third. And the metal/metal oxide chemistry is handled in an uncoupled manner in the fourth.

## a. Decay of Oxygen Excited States

Reactions 33 and 34 involve the decay of  $[O(^1S)]$  and  $[O(^1D)]$ .

The  $O(^1S)$  decay is simply expressed as

$$[O(^1S)] = [O(^1S)]_0 e^{-k_{33}t} \quad (63)$$

and the differential equation for  $O(^1D)$

$$\frac{d[O(^1D)]}{dt} = -k_{34}[O(^1D)] - d[O(^1S)]/dt \quad (64)$$

has the solution

$$[O(^1D)] = \left\{ [O(^1D)]_0 - k_{33}[O(^1S)]_0 / (k_{34} - k_{33}) \right\} e^{-k_{34}t} + k_{33}[O(^1S)]_0 / (k_{34} - k_{33}) e^{-k_{33}t} \quad (65)$$

where the initial values  $[O(^1D)]_0$  and  $[O(^1S)]_0$  are those updated at the end of the charged species treatment. At the same time,  $[O_2(^3\Sigma)]$  is decayed according to Reaction 23

$$[O_2(^3\Sigma)] = [O_2(^3\Sigma)]_0 e^{-k_{23}[O_2]t} \quad (66)$$

#### b. Nitrogen-Nitric Oxide

The interplay of [N] and [NO] is handled by dividing the solution for [N] into two parts separated by an updating of [NO].

The first set of nitrogen Reactions 1, 2, and 32 are faster and do not involve [NO].  $N(^2P)$  decays in a simple exponential manner according to Reactions 1 and 32, in analogy to  $O(^1S)$  in Eq (63). The differential equation for  $N(^2D)$ , derived from Reactions 2 and 32, is solved as for  $O(^1D)$  and the solution is of the same form as Eq (65). And  $N(^4S)$  is exponentially decayed according to Reaction 1.

For the updating of [NO], the pertinent reactions of Table II are divided into two groups: Reactions 1, 2, 3, 4 and Reactions 24, 25, 26, 28, and 31. The solution to the first group of four reactions is found in Eq (57) and Eq (58) as used in Section IV-7. The second group of five are

divided into a formation term,  $F = k_{26} [O][NO_2] + k_{31}[N(^4S)][O_3]$ , and a destruction term,  $D = k_{24}[O_3] + k_{25}[O] + k_{28}[N]$ , for  $[NO]$  and FUNCTION WQX is used to evaluate  $[NO]_t$ . Once both solutions have been found, the maximum of the two is adopted as the final value of  $[NO]$  at  $t_c$ .

With  $[NO]$  updated, the second set of nitrogen Reactions 3, 4, 27, 28, 29, and 31 are considered. Each atomic state of nitrogen suffers a simple exponential decay with exponents derived from: Reactions 3, 27, 28, and 29 for  $N(^2P)$ , Reactions 4, 27, 28, and 29 for  $N(^2D)$ , and Reactions 3, 27, 28, 29, and 31 for  $N(^4S)$ . The total change in  $[N]$  over the two part treatment is calculated and added to  $[O]$  since the major reactions in the nitrogen treatment involve production of atomic oxygen.

### C. Ozone and Atomic Oxygen

Ozone is involved in Reactions 5, 6, 21, 23, 24, 30, 31, and photo dissociation. Neglecting Reactions 21 and 30, the differential equation for  $[O_3]$  may be written

$$d[O_3]/dt = - [O_3] \left\{ k_6[O] + k_{24}[NO] + k_{31}[N(^4S)] + \phi \right\} + k_5[O_2][O] + k_{23}[O_2][O_2(^3\Sigma)] \quad (67)$$

where  $\phi$  is the photodissociation rate constant for ozone and Eq(66) can be substituted for  $[O_2(^3\Sigma)]$ . The value of  $[O]$  in the second term is replaced by  $[O]e^{-k_5[O_2]t}$  and Eq (67) is solved to give the final value for  $[O_3]$

$$[O_3] = \left\{ [O_3]_o + \gamma[O_2(^3\Sigma)]_o/(\gamma-\beta) + \delta[O]/(\delta-\beta) \right\} e^{-\beta t} - \left\{ \gamma[O_2(^3\Sigma)]_o/(\gamma-\beta) \right\} e^{-\gamma t} - \left\{ \delta[O]/(\delta-\beta) \right\} e^{-\delta t} \quad (68)$$

where  $\beta = k_6[O] + k_{24}[NO] + k_{31}[N(^4S)] + \phi$ ,  $\gamma = k_{23}[O_2]$ , and  $\delta = k_5[O_2]$ .

With Reactions 33 and 34 already addressed, atomic oxygen is involved in Reactions 5, 6, 23, 25, 26, and 29 for which the differential equation is

$$\frac{d[O]}{dt} = - [O] \left\{ k_5[O_2] + k_6[O_3] + k_{25}[NO] + k_{26}[NO_2] \right\}$$

$$+ \left\{ k_{29}[NO_2][N] + \phi \right\} + k_{23}[O_2][O_2(^3\Sigma)] \quad (69)$$

If  $[O_2(^3\Sigma)]$  is replaced by Eq (66), the solution to Eq (69) using the updated value for  $[O_3]$  is

$$[O] = \left\{ [O]_0 + \gamma [O_2(^3\Sigma)]_0 / (\gamma - \beta) \right\} e^{-\beta t} + \delta / \beta (1 - e^{-\beta t}) - \left\{ \gamma [O_2(^3\Sigma)]_0 / (\gamma - \beta) \right\} e^{-\gamma t} \quad (70)$$

where  $\beta = k_5[O_2] + k_6[O_3] + k_{25}[NO] + k_{26}[NO_2]$ ,  $\gamma = k_{23}[O_2]$ , and  $\delta = k_{29}[NO_2][N] + \phi$ . The above solution is applied to  $[O(^3P)]$  and the values of  $[O(^1S)]$  and  $[O(^1D)]$  are adjusted in proportion to the resultant change in  $[O(^3P)]$ .

#### d. Inclusion of Continuous Source Production

The second step of the neutral chemistry includes continuous energy deposition contributions in the species time dependent solutions. The species and their reactions from step one are addressed again here. For each species, formation rates ( $\text{sec}^{-1}$ ) and destruction rates ( $\text{cc}^{-1}\text{sec}^{-1}$ ) are calculated using the continuous source rates from ENQDEP and the appropriate reaction rate constants and current species concentration values. Then, in sequence, the solutions for  $O(^3P)$ ,  $O(^1D)$ ,  $O(^1S)$ ,  $N(^4S)$ ,  $N(^2D)$ ,  $N(^2P)$ ,  $O_2(^3\Sigma)$ , and  $O_3$  are systematically calculated using FUNCTION WQX, with the calculational time,  $t_c$ , and concentrations at  $t = 0$  assumed to be zero. The final value for each species concentration at  $t_c$  is computed as the sum of the solutions from the first and second steps of the neutral species chemistry.

e.  $\text{NO}_2$ ,  $\text{N}_2\text{O}$ , and  $\text{O}_2(^1\Delta)$ 

The  $\text{NO}_2$  and  $\text{N}_2\text{O}$  concentrations are calculated using an approach to equilibrium with the FUNCTION WQX.  $\text{NO}_2$  is formed in Reactions 24 and 25 and destroyed in Reaction 26, 27, and 29.  $\text{N}_2\text{O}$  is formed in Reactions 28 and 29 and destroyed in Reaction 30. Appropriate formation and destruction terms are computed for each species and FUNCTION WQX is applied to calculate their concentrations at  $t_c$ .

Having updated  $[\text{N}_2\text{O}]$ ,  $\text{O}_2(^1\Delta)$  is now addressed. It's only source of production is from the continuous energy deposition and its only form of destruction is by way of Reaction 30. FUNCTION WQX is utilized twice to solve for  $[\text{O}_2(^1\Delta)]$  and the final value of the concentration is taken as the sum of the two results. In the first, both production and destruction terms are considered and  $[\text{O}_2(^1\Delta)]_0$  is set equal to zero; and in the second, the production term is deleted and  $[\text{O}_2(^1\Delta)]_0$  is its initial value.

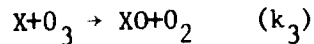
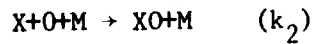
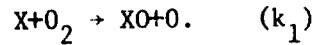
## f. Hydrogen

The final neutral system to be addressed before the metal oxides is the H, OH equilibrium of Reactions 21 and 22. Again the FUNCTION WQX is employed in an ad hoc fashion to solve this cyclic set of equations. In this case, the value of [OH] at  $t_c$ , is calculated using the production term  $k_{21}[\text{O}_3] \{[\text{H}] + [\text{OH}]\}$ , the destruction term  $k_{21}[\text{O}_3] + k_{22}[\text{O}(^3\text{P})]$ , and the initial value  $[\text{OH}]_0$ . The value of [H] is then taken as that free hydrogen which has not been converted to OH.

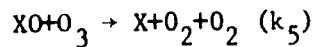
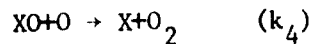
## g. Metal Oxides

The metal and metal oxide reactions are uncoupled from the rest of the analytical algorithm. The light metals such as Li, Be, Mg, Al, Si, Ti, Fe, and Cu, presently carried in WORRY, are treated in the same manner; therefore, in the following discussion the symbol X will be used to describe

any one of these eight metals. The reactions and rate constants for the formation of the metal oxides are



The reactions and rate constants for the destruction of the metal oxides are



The differential equation for  $[XO]$  is

$$\frac{\partial [XO]}{\partial t} = \left\{ k_1[O_2] + k_2[O][M] + k_3[O_3] \right\} [X] - [XO] \left\{ k_4[O] + k_5[O_3] \right\} \quad (71)$$

But the total number of metal atoms does not change

$$[X] + [XO] = A \quad (72)$$

and  $A$  is a constant, whose value is  $[X]_0$  if all metal atoms are unoxidized at  $t = 0$ .

Thus Eq (71) can be rewritten as

$$\frac{d[XO]}{dt} = F(A-[XO]) - D[XO] \quad (73)$$

where  $F = k_1[O_2] + k_2[O][M] + k_3[O_3]$  and  $D = k_4[O] + k_5[O_3]$ .  $F$  and  $D$  are computed for each species using the final values of  $[O]$ ,  $[O_2]$ , and  $[O_3]$ , which are assumed to be constant.

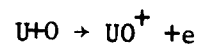
Reordering Eq (73) gives



$$\frac{d[XO]}{dt} = FA - (F+D)[XO] \quad (74)$$

which may be solved for  $[XO]$  at  $t_c$  using FUNCTION WQX with a formation term of FA, a destruction term of  $F + D$ , and  $[XO]_0$ .

Because of the unique chemistry of U, only the following reaction is treated and in a simple exponential fashion:



With all species concentrations now calculated and stored in the CMD array, the analytic algorithm is complete and new formation and destruction rates for all species are calculated where needed, to be used in subroutine RRCRT.

## SECTION V

## ACCESSORIAL SUBROUTINES AND FUNCTIONS

## 1. ROUTINES SHARED BY BOTH ALGORITHMS

## a. ENTDEP and ENQDEP

The routine ENTDEP calculates the initial distribution of ions and neutrals produced at  $t = 0$  by the prompt energy deposition, while its entry point ENQDEP partitions the continuous energy deposition (ion pairs/cc/sec) among the same set of species. ENTDEP also calculates the specific heat of the sample volume, the internal kinetic energy of the gas, and adjusts the gas temperature accordingly.

ENTDEP requires inputs of the gas temperature, the total density, and the total number of ion pairs produced by the prompt energy deposition. The latter two are converted to a fractional ionization and ionized fractions of  $[N_2]$ ,  $[O_2]$ , and  $[O]$  are calculated. Then the concentration for each of the following species is adjusted for its production or depletion by ionization of  $N_2$ ,  $O_2$ , and  $O$ :  $N_2$ ,  $O_2$ ,  $O(^3P)$ ,  $O(^1D)$ ,  $O(^1S)$ ,  $N(^4S)$ ,  $N(^2D)$ ,  $N(^2P)$ ,  $O_2(^1\Delta)$ ,  $O_2(^3\Sigma)$ ,  $N^+$ ,  $O^+$ ,  $N_2^+$ ,  $O_2^+$ , and  $e^-$ .

ENQDEP likewise calculates the continuous ionization rates of  $N_2$ ,  $O_2$ , and  $O$  and converts these to formation and destruction rates for the same set of product species.

## b. RATE

The RATE routine simply calculates the rate constants for the reactions of table II and stores them in the R array. RATE requires inputs of the gas temperature ( $^{\circ}K$ ) and the total density ( $cc^{-1}$ ). All three body reactions are provided with pseudo-second-order rate constants by including  $[M]$  in the constant.

## c. RRCRT

This routine stores the necessary species destruction rates in RCRT and transforms the species formation rates into vibrational level formation rates which are stored in the RR array and used by the WORRY calculation of chemiluminescence. Because the product vibrational level branching ratios for the various reactions are not known, equal partition to the highest accessible level is presently assumed.

## 2. FUNCTIONS OF THE ANALYTIC ALGORITHM

## a. FUNCTION WQX (A, B, XO, T)

This function solves the differential equation

$$\frac{d[X]}{dt} = A - B[X] \quad (75)$$

to give [x] at time, T, with  $[x]_0 = X0$ .

## b. FUNCTION WR(A, B, XO, T)

This function solves the differential equation

$$\frac{d[x]}{dt} = A - B[x]^2 \quad (76)$$

to give [x] at time, T, with  $[x]_0 = X0$ .

## c. FUNCTION SHXOX (X)

This function evaluates the expression

$$\frac{e^x - 1}{x}$$

## d. FUNCTION XIIGL (A, B, C, T)

This function evaluates the integral

$$\int_0^T \frac{e^{-Bt} dt}{(1 + ACt)}$$

e. FUNCTION X2IGL (A, B, C, T)

This function evaluates the integral

$$\int_0^T \frac{e^{-Bt} dt}{(1 + ACt)^2}$$

i. FUNCTION EIF (X)

This function evaluates the integral

$$\int_X^\infty \frac{e^{-x}}{x} dx$$

SECTION VI  
COMPARISONS OF ALGORITHMS

Figures 2, 3, 4, and 5 demonstrate the agreement between the analytic and numerical algorithms. A large number of such comparisons were presented at the DNA 1973 Atmospheric Effects Symposium at San Diego (ref 1). The problem definitions for figures 2, 3, and 4 were taken from Ory and Gilmore (ref 2), while that for figure 5 was taken from Scheibe (ref 3).

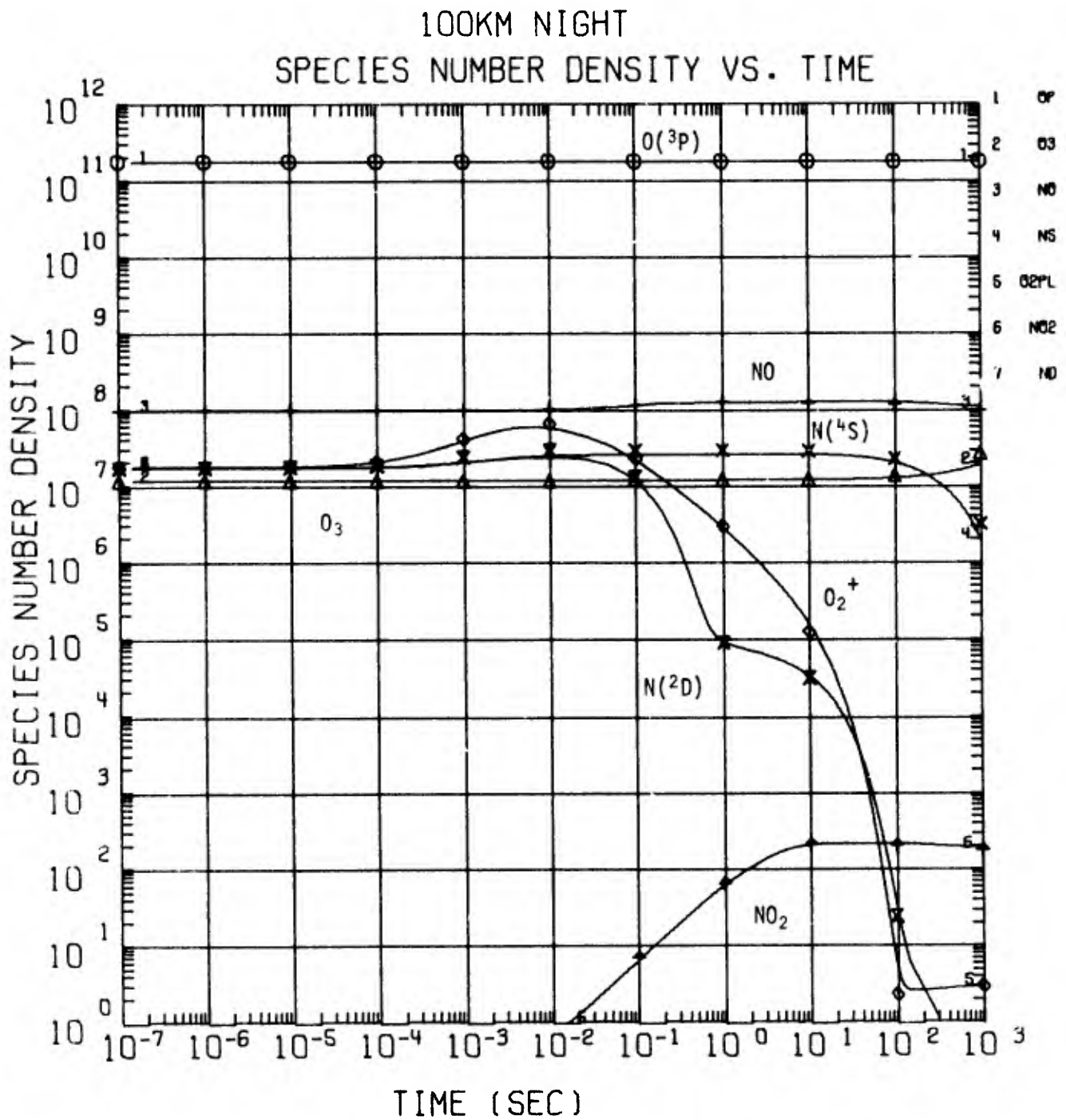


Figure 2. Comparison of Numerical (—) and Analytic (symbols) Algorithms (Altitude of 100 km, energy deposition of  $10^8$  ion pairs/cc during the night).

60KM NIGHT  
SPECIES NUMBER DENSITY VS. TIME

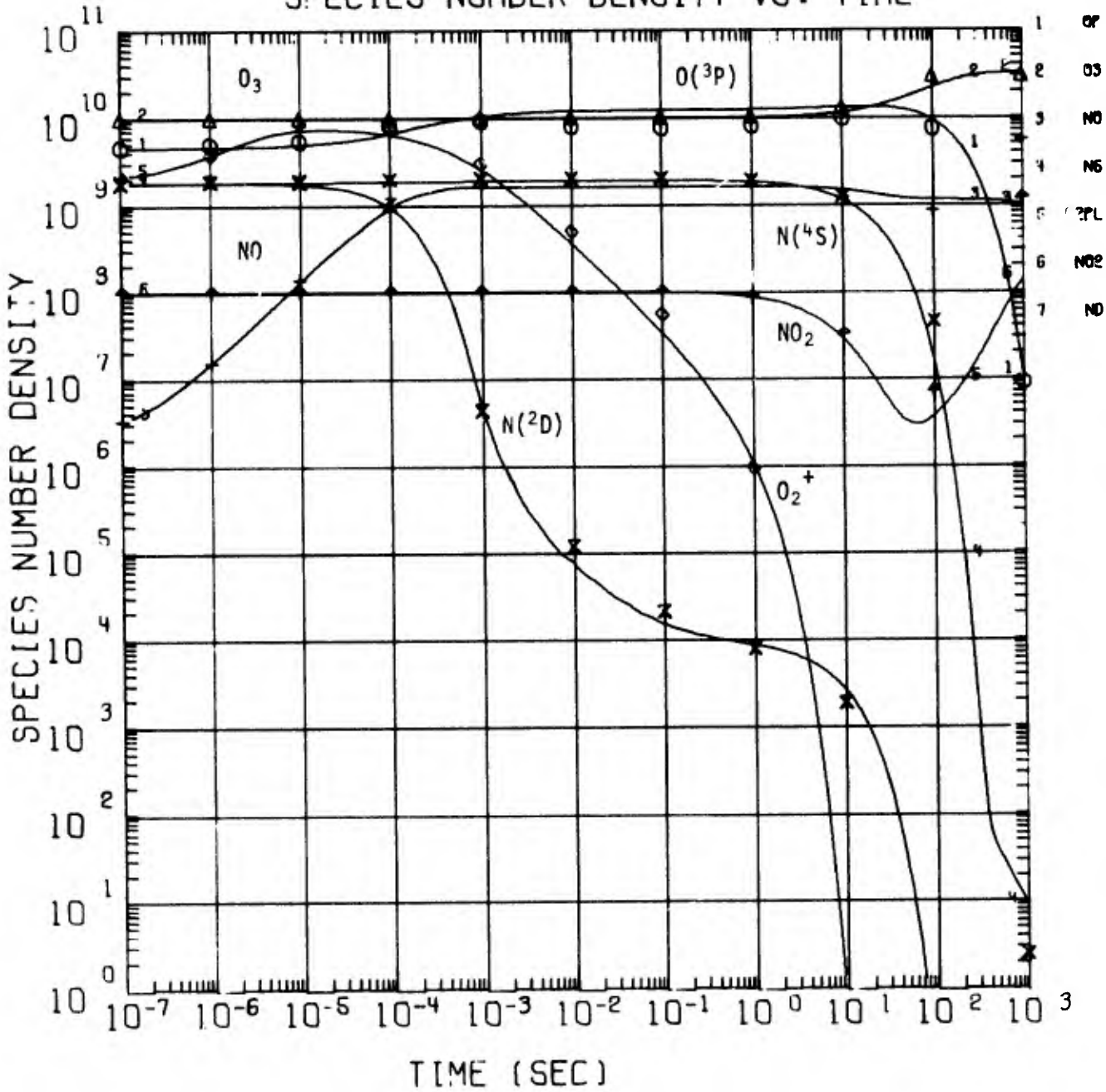


Figure 3. Comparison of Numerical and Analytic Algorithms (Altitude of 60 km, energy deposition of  $10^{10}$  ion pairs/cc during the night).

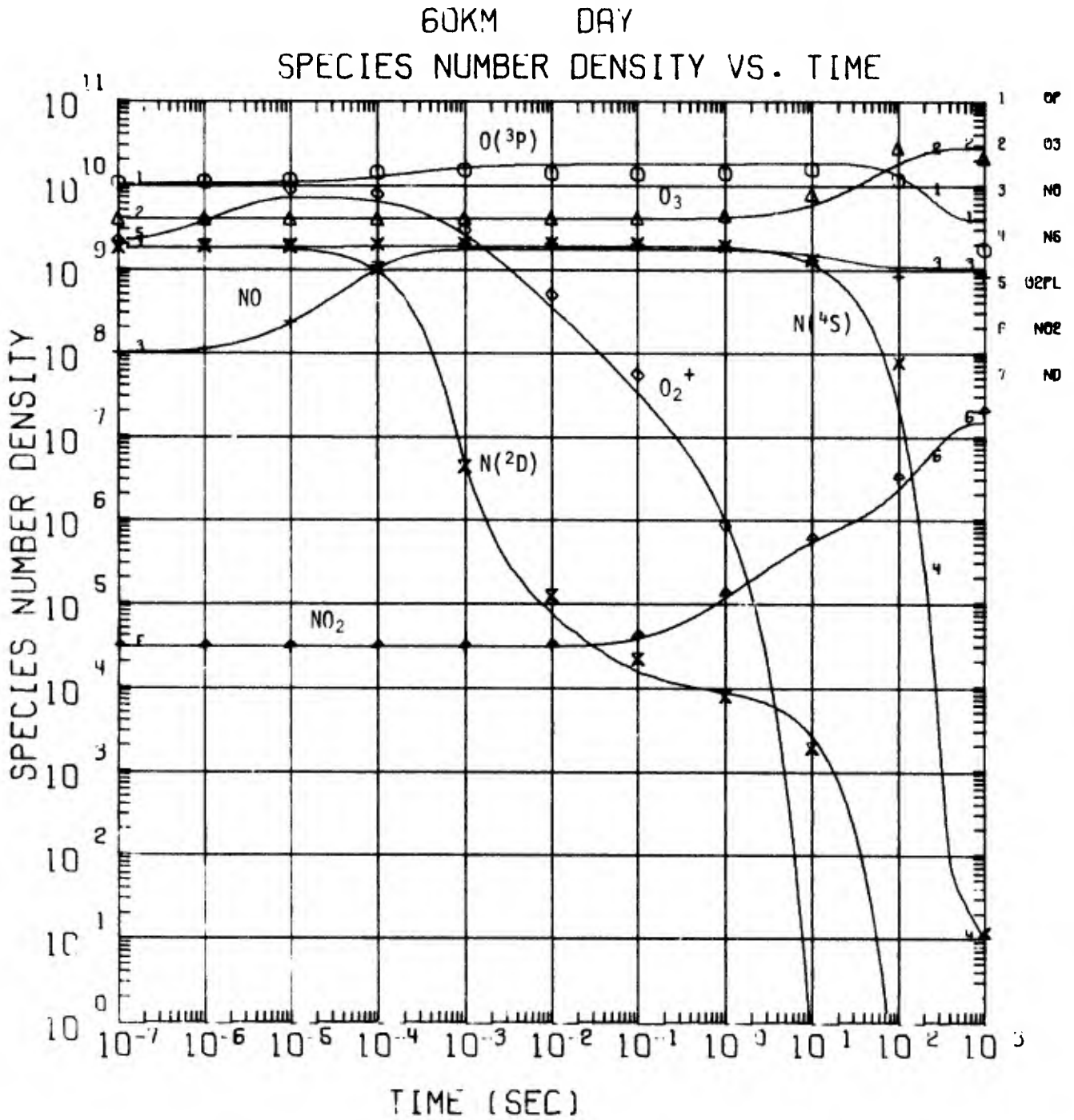


Figure 4. Comparison of Numerical and Analytic Algorithms (Altitude of 60 km, energy deposition of  $10^{10}$  ion pairs/cc during the day).



40KM NIGHT  
SPECIES NUMBER DENSITY VS. TIME

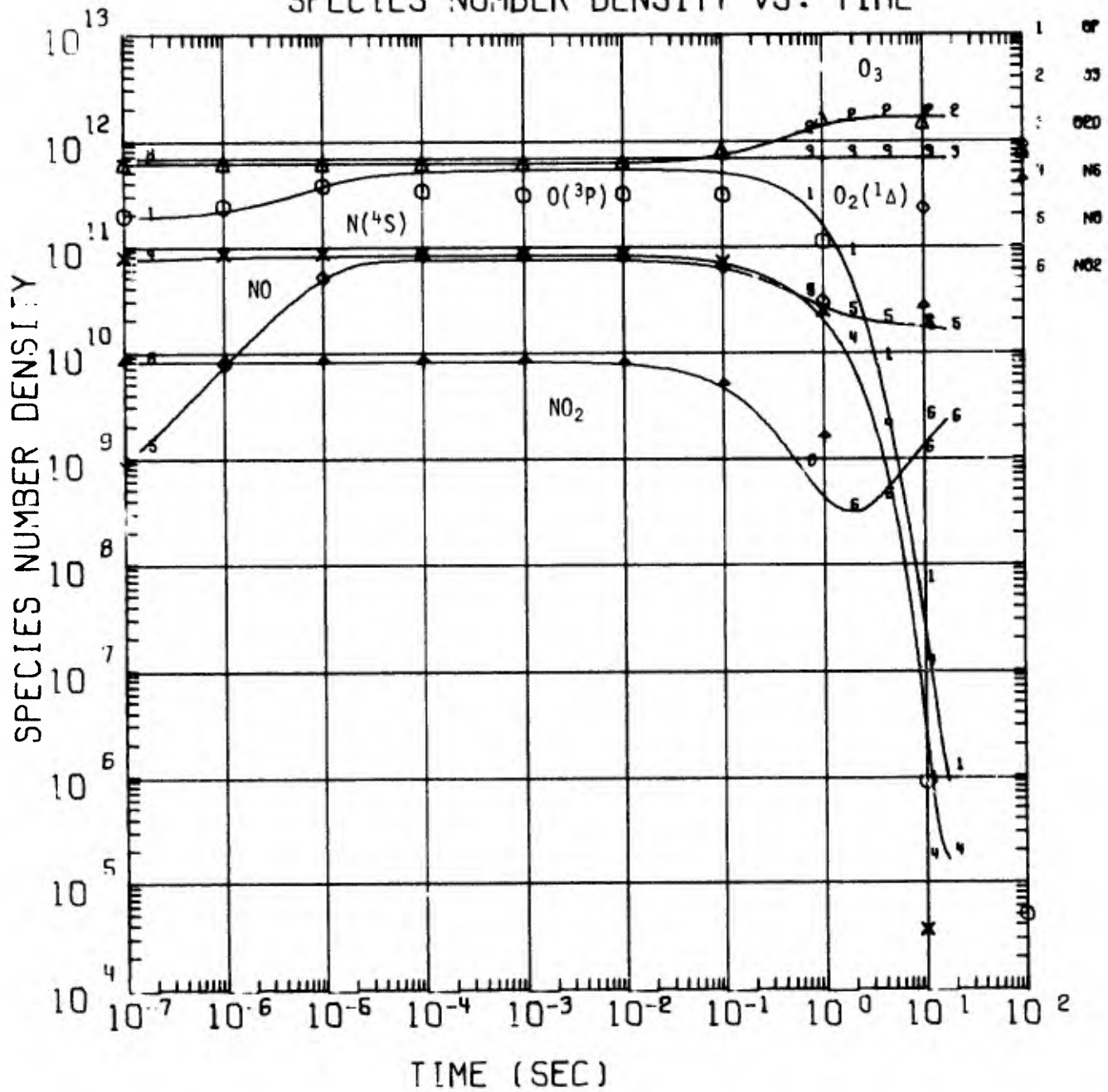


Figure 5. Comparison of Numerical and Analytic Algorithms (Altitude of 40 km, Energy Deposition of  $4 \times 10^{11}$  ion pairs/cc, and continuous ionization of  $10^7$  ion pairs/cc/sec during the night).

SECTION VII  
CODE LISTING

This section lists the chemistry package as it is run outside WORRY. The first two routines constitute the external diver which decodes the users problem and prepares the necessary inputs for a chemistry package calculation by either the analytic or numerical algorithms. The diver also uses two plotting routines (not listed) to prepare comparison plots between the two methods. Since the external running of the chemistry package requires diagnostics which are otherwise unnecessary, the routines as listed contain some statements which are not standard in the operational version of WORRY. These statements are identified by a CH prefix in columns 73 and 74 of the card containing the statement.

When the chemistry package is used within WORRY, explicit phenomenology calculations provide a description of the continuous energy deposition as a function of time. When running in an external mode, the chemistry package as listed must assume a form for the time dependence of the continuous energy deposition. That form is

$$Q = QI*(1 + t)** 1.2$$

where  $t$  is the time after prompt energy deposition and  $QI$  is the energy deposition rate at  $t = 0$ .





LINE NO	DESCRIPTION	J.	NR	TIME	DEP	60KM	NIGHT	VERSION	Y	NO2	NO	YPF C
143	CONTROL TM	1.0E-7		1.99E12		60KM	NIGHT	1.0E-4	0.			
144	CONTROL TM	1.0E-7		1.0E-6		100.	100.	1.0E-3	1.0E-2	1.0E-2	1.0E-	
145	CONTROL TM	60		260								
146	CONTROL TM	.535										
147	CONTROL TM	0.		0.								
148	CONTROL TM	0.		0.								
149	CONTROL TM	0.		0.								
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214	CONTROL TM	0.		0.								
215	CONTROL TM	0.		0.								
216	CONTROL TM	0.		0.								

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*****
CONTINUE
KEMWIND 4
KEMWIND 1
KEMWIND 9
DO 2 I=1,15
YDMETH(I)=0.
JYMETL(I)=3.
DO 3 I=1,70
I(I)=0.
DO 4 I=1,NSPEC
XMD(I)=3.
IPLM=0
IIMH=3
PRINT 30
CONTINUE 2
READ (3,37) ALPHA
WRITE (6,35) EQ.1MS) JO TO 7
IF (ALPHA(1).EQ.1MS) GO TO 7
WRITE (2,37) ALPHA
IF (ALPHA(1).EQ.1MT) GO TO 6
IF (ALPHA(1).EQ.1MP) GO TO 27
READ (2,37) TYPE,(X(L),L=1,7)
JUMP=LOGIC(3,TYPE,TEST)
GO TO (5,7,9,11,12,13,15,17,19,14), JUMP
CONTINUE
*****
TITLE FOR PLOTS--USES ONLY COLUMNS 10 TO 60
REWIND 2
READ (2,43) TYPE,(ITL(L),L=1,7)
GO TO 5
CONTINUE
SWITCHES
DO 8 I=1,60
SWTCH(I)=ALPH(I)
GO TO 5
CONTINUE
CONTROL TH TIME CARDS
DO 10 I=1,7
ITL(I)=ITM
ITM=ITM+7
GO TO 5
CONTINUE
CHD CONCENTRATIONS OF N2,O2,O,(ANG M)
CN2=X(1)
CO2=X(2)
CO=X(3)
CNS=X(4)
GO TO 5
CONTINUE
ALTITUDE AND TEMPERATURE
AL=X(1)
TEM=X(2)
GO TO 5
CONTINUE
ENERGY DEPOSITION ED,OI,UP,QA
ED=X(1)
JI=X(2)
DP=X(3)

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6225 EXPB0T=EXP(-B0*DT)
6226 COPL0=COPL
6227 GM=RNPL0*CO
6228 YAE=GM*CNPL/(B0-BN)
6229 COPLT=(COPLG-YAE)*EXPB0T+YAE*EXPB0T
6230 COPL0=MAX(COPLT,0.0,DT)
6231 COPL=COPLT+COPLT
6232
6233 --- ATOMIC IONS
6234 ATPLT=COPLT+CNPLT
6235 X0GL=(1.-EXP(-B0*DT))/B0
6236 COPL01=COPL0-YAE
6237 COPL02=CNPL01+YAE
6238 X0IGL=COPL01+X0GL+YAE*XNIGL
6239 X0GLP=COPL01+X0GL+COPL02*XNIGL
6240 X0GLP2=X0IGL*CO2+X0PL02
6241
6242 --- NEGATIVE CHARGE
6243 ALPHAR=(RN2PEL*CN2PL+R02PEL*CO2PL+RNOPEL*CN0PL)/(CN2PL+CO2PL+CN0PL
6244 +COPL+CNPL)
6245 AB=ALPHAR*ATPLT*DT
6246 ALPHAR=ALPHAR*SMXOX(AB)
6247 CNEG=CEL+CO2NG
6248 CNEG0=CNEG
6249 ANG0P=ALPHAR*CN0EG
6250 ANG0PT=1.+ANG0*DT
6251 CNEG=ATPLT+(CNEG0-ATPLT)/ANG0PT
6252 CNEG0=MR(CQ,ALPHAR,JUC,DT)
6253 CNEG=AMAX(CNEG,CNEG0)
6254 X0GLP1=X1IGL(ALPHAR,BN,CNEG,DT)
6255 X0GLP2=X1IGL(ALPHAR,BN,CNEG,DT)
6256 X0GLP=COPL01+X0GLP1+COPL02+X0GLP2
6257
6258 --- N2+
6259 CN2PL0=CN2PL
6260
6261 --- FACTOR OF 2 DUE TO SEPARATION OF REACTION 9 INTO TWO PARTS
6262 BN2=AN2P02*CO2+XN2PNO*CN0+RN2PEL*CO*2.0
6263 CN2PLT=CN2PL0*EXP(-BN2*DT-RN2PEL*(X0GLP-X0GLP))/(1.+ALPHAR*CN0EGJ*0
6264 IT)+1.E-50
6265 BN2EFF=BN2+(RN2PEL*(X0GLP-X0GLP))/DT
6266 IF (CN2PLT.LE.2.E-50) BN2EFF=BN2
6267 GAMMA=SQRT(ALPHAR*ALPHAR)
6268 XN2IGL=X1IGL(ALPHAR,BN2EFF,CNEG,DT)
6269 B=RN2PEL+CNEG+BN2
6270 CN2PL0=MAX(CN2PLT,B,0,DT)
6271 CN2PL=ARMAX(CN2PL0,CN2PL0)
6272 XN2P02=CN2PL0+XN2IGL*CN2P02*CO2
6273 XN2PNO=CN2PL0+XN2IGL*CN2PNO*CN0
6274 DELTA1=BN2EFF+BN
6275 DELTA2=BN2EFF+BN
6276 XN2PEL=RN2PEL*CN2PL0*(COPL01+X1IGL(ALPHAR,DELTA1,CNEG,DT)+COPL02*
6277 X1IGL(ALPHAR,DELTA2,CNEG,DT)+CNEG0*X2IGL(GAMMA,BN2EFF,CNEG,DT))-C
6278 COPL01*X2IGL(GAMMA,DELTA1,CNEG,DT)-COPL02*X2IGL(GAMMA,DELTA2,CNEG,
6279 DT)
6280 XN2=GN2P02+GN2PNO+GN2PL0+GN2PEL+1.E-30
6281 DELCN2P=CN2PL0-CN2PLT
6282 DELCN2R=DELN2P/ADELN2
6283
6284 --- UPDATE N ANU 0
6285 GOPLN2=RN2OPL+CN2*X0IGL
6286 GNPL02=XCN2NPL+CNPL0*CO2+XNIGL
6287 GNPL02=RO2NPL+CO2+XNIGL+CNPL0
6288 GNPLNO=GM*CNPL0+XNIGL
6289 GNPLNO=XCNPLNO+XNIGL
6290 GNPLNO=XCNPLNO+XNIGL
6291

```





```

0000 --- 0 + UC + : U3 + M
0000 --- B1=I020*C02
0000 --- 02 + 02*(S) 03 + 0
0000 --- B11=R02502*C02
0000 --- 0(1S) 0(10)+ HV
0000 --- B13=R0SHV 0(10) HV
0000 --- B14=R0DHV 0(10)+ HV
0000 --- N(4S)+ 03 NO** + 02
0000 --- B16=RNS03*C03
0000 --- DELAY 0(1S) AND 0(10) AND 02*(S)
0000 --- XCOSD=COS*EXP(-B13*DT)
0000 --- COS=C00+B13/(B13-B14)*COS)*EXP(-B14*DT)-B13/(B13-B14)*XCOSD
0000 --- COP=C00-COS
0000 --- COZS=C02S9*EXP(-B11*DT)
0000 --- N(4S)+ 02
0000 --- N(2P)+ 02 NO** + 0
0000 --- B3=RNS02*C02 NO** + 0
0000 --- N(20)+ 02 NO** + 0
0000 --- B4=RND02*C02 N(20)+ HV
0000 --- N(2P)
0000 --- B5=RMPHV
0000 --- ONP=CNP
0000 --- OND=CND
0000 --- ONS=CNS
0000 --- DECAY STATES OF N
0000 --- CNP=CNP*EXP(-(B5+B3)*DT)
0000 --- CND=(CND+B3/(B5-B3)*ONP)*EXP(-B4*DT)-B5/(B5-B3)*CNP
0000 --- CNS=CNS*EXP(-B3*DT)
0000 --- CN=CNP+CND+CNS
0000 --- RECALCULATION OF NO
0000 --- BALL=R03NO*L03+I0NO*CO+I0NO*CN
0000 --- CALL=B19*CNS+R0NO2*CO*CN02
0000 --- CNO=CNO*(CALL,BALL,CNO2,DT)
0000 --- XKAY=R0ND02+RNDNO
0000 --- B1A=R04+RNDNO*CN0
0000 --- B1A=RTA*DT
0000 --- B1O=C02NO2+RND02/XKAY
0000 --- CMAE=RTO-(RTU-CNO2)*EXP(-XKAY*ZND*DT+SHXOX(B1A))
0000 --- CMAE=RTA*DT
0000 --- CMAE=RTA*DT
0000 --- ALLD=RNS02+RNSNO
0000 --- RTA=RNS02+C02NO2/ALLD
0000 --- CNG=RTI-(RTI-CNO)*EXP(-ALLMU*(ZNS+ZNP)*DT+SHXOX(SHAT))
0000 --- CNU=ARHAI(CNU,CNO2)
0000 --- DECAY OF STATES OF N
0000 --- BALL=(RNO2NA+KNO2N3)*CNO2+TNON*CN0
0000 --- B3=RNS03+CNO

```







```

1074 FN02=C1+0+C141
1075 DN02=B145*B111+B112
1076 CN02=MOX(FM02, DN02, CN02, DT)
1077 C111=B111*CN02
1078 FN20=C113+C111
1079 DN20=MQA(FN20, DN20, CN20, DT)
1080
1081 --- N20 + O2(10) N2 + O3
1082
1083
1084 B145=R02N20*CN20
1085
1086 --- CALCULATION OF O2(10)
1087
1088 CO20Q=MQX(Q020, B145, O, C, DT)
1089 CO20=MQX(C, B145, CO20, DT)
1090 CO20=CO20Q+CO20
1091
1092 --- H + O3 OH + O2
1093
1094
1095 B21=RMO3*CO3
1096
1097 --- OH + O OH + O2
1098
1099 B22=ROMO*CO
1100
1101 --- CALCULATION OF H AND OH
1102
1103 QA=B21
1104 QY=B22
1105 CHOH=CH+COH
1106 QZ=CHOH*QX
1107 COH=MOX(QZ, (QA+QY), COM, DT)
1108 CH=MAX1((CHOH-COH), 0.)
1109
1110 --- BEGIN METAL-METAL OXIDE CALCULATION
1111 IF (CMTL.LE.0.) GO TO 17
1112
1113 --- X + O3 XO + O2
1114
1115 FMTLOZ=RMTLO3*CO3
1116
1117 --- LITHIUM
1118
1119 NMT=1
1120 IF (GLI.LE.0.) GO TO 6
1121
1122 --- LI + O + M LIO + M
1123
1124 B211=TLIO*CO2
1125 C211=9211*GLI
1126
1127 --- LI + O2 LIO + O
1128
1129 B212=8L102*CO2
1130 C212=8212*GLI
1131
1132 --- LIO + O LIO + O2
1133
1134 B213=R0L10*CO2
1135
1136 --- LIO + O3 LIO + O2 + O2
1137
1138 B214=R0JL10*CO3
1139
1140 --- LI + O3 LIO + O2
1141
1142 C215=FMTLOZ*GLI
1143 QY=B211+B212+FMTLOZ
1144 QV=B213+B214
1145 GO TO 15
1146 CONTINUE
1147
1148

```

```

1151 3CRYLLIUM
1152 NMT=NMT+1
1153 IF (CBE.LE.0.) GO TO 3
1154 BE + 0 + M
1155
1156 C221=B211*CBE
1157 BE + 0
1158
1159 B222=R3E02*CO2
1160 C222=B222*CBE
1161
1162 BE0 + 0
1163 BE0 + 02
1164
1165 BE0 + 03
1166 B224=R03610*CO3
1167 BE + 03
1168
1169 C245=FMLOZ*CBE
1170 QX=B211+B222+FMLOZ
1171 QY=B213+B224
1172 GO TO 15
1173 CONTINUE
1174
1175 MAGNESIUM
1176 NMT=NMT+1
1177 IF (CMG.LE.0.) GO TO 14
1178 MG + 0 + M
1179
1180 C231=B211*CMG
1181 MG + 02
1182
1183 B232=RMGO2*CO2
1184 C232=B232*CMG
1185
1186 MGO + 0
1187 MGO + 03
1188
1189 B234=R03MG0*CO3
1190 MG + 03
1191
1192 C235=FMLOZ*CMG
1193 QX=B211+B232+FMLOZ
1194 QY=B213+B234
1195 GO TO 15
1196 CONTINUE
1197
1198 ALUMINUM
1199 NMT=NMT+1
1200 IF (CAL.LE.0.) GO TO 14
1201 AL + 0 + M
1202
1203 C241=B211*CAL
1204 AL + 02
1205
1206 B242=RAL02*CO2
1207 C242=B242*CAL
1208
1209 AL0 + 0
1210 B243=R0AL0*CO3
1211 AL + 03
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1299 0 --- FE + 0 + M          FE0 + M          027
1300 0 --- G271=B211*CFE          FE0 + 0          023
1301 0 --- FE + 02          FE + 02          022
1302 0 --- B272=RF502*CO2          FE + 02          023
1303 0 --- C272=B272*CFE          FE + 02          023
1304 0 --- FE0 + 0          FE0 + 0          023
1305 0 --- FE0 + 03          FE0 + 03          023
1306 0 --- B274=R03FE0*CO3          FE0 + 02          023
1307 0 --- FE + 03          FE0 + 02          023
1308 0 --- G275=FMLOZ*CFE          CU0 + M          023
1309 0 --- QX=B211+B272+FMLOZ          CU0 + 0          023
1310 0 --- QY=B211+B274          CU0 + 0          023
1311 0 --- GO TO 15          CU0 + 0          023
1312 0 --- CONTINUE          CU0 + 0          023
1313 0 --- COPPER          CU0 + 0          023
1314 0 --- NMT=NMT+1          CU0 + 0          023
1315 0 --- IF (CCU.LE.0.) GO TO 16          CU0 + 0          023
1316 0 --- CU + 0 + M          CU0 + 0          023
1317 0 --- C261=B211*CCU          CU0 + 0          023
1318 0 --- CU + 02          CU0 + 0          023
1319 0 --- B262=RCU02*CO2          CU0 + 0          023
1320 0 --- C262=B232*CCU          CU0 + 0          023
1321 0 --- CU0 + 0          CU0 + 0          023
1322 0 --- CU0 + 03          CU0 + 0          023
1323 0 --- B264=R03CU0*CO3          CU0 + 0          023
1324 0 --- CU + 03          CU0 + 0          023
1325 0 --- C265=FMLOZ*CCU          CU0 + 0          023
1326 0 --- QX=B211+B262+FMLOZ          CU0 + 0          023
1327 0 --- QY=B213+B284          CU0 + 0          023
1328 0 --- CONTINUE          CU0 + 0          023
1329 0 --- NH=NMT+1          CU0 + 0          023
1330 0 --- QZ=CMD(NM)+CMD(NM+10)          CU0 + 0          023
1331 0 --- QZ=CMD(NM)+CMD(NM+10),OT          CU0 + 0          023
1332 0 --- QZ=CMD(NM)+CMD(NM+10),C.U          CU0 + 0          023
1333 0 --- GO TO (8,9,10,11,12,13,14,15), NMT          CU0 + 0          023
1334 0 --- CONTINUE          CU0 + 0          023
1335 0 --- CONTINUE          CU0 + 0          023
1336 0 --- CALCULATION OF U0+          CU0 + 0          023
1337 0 --- IF (CU.LE.5.) GO TO 18          CU0 + 0          023
1338 0 --- U + 0          CU0 + 0          023
1339 0 --- B214=RU0*CCU          CU0 + 0          023
1340 0 --- XCUD=CU          CU0 + 0          023
1341 0 --- CU=CU*EXP(-B210*OT)          CU0 + 0          023
1342 0 --- CUOPL=CUOPL+(XCUD-CU)          CU0 + 0          023
1343 0 --- CONTINUE          CU0 + 0          023
1344 0 --- RECALCULATE NECESSARY FORMATION AND DESTRUCTION RATES          CU0 + 0          023
1345 0 --- CMI=CLI+CBE+CMG+CAL+CSI+CTI+CFE+CCU          CU0 + 0          023
1346 0 --- CMILO=CLIO+CBEO+CMGO+CALO+CSIO+CFEO+CCUO+CTIO          CU0 + 0          023
1347 0 --- C1=BI*CO          CU0 + 0          023

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C --- FUNCTION X2IGL (A,B,C,T)
C --- INTEGRAL OF EXP(-B*T)*DT/(1.+A*C*T)**2
      CC=A*C
      AA=B/CC
      IF (AA.GT.700.0) GO TO 1
      BB=A*(1.+CC*T)
      XIGL=AA*EXP(AA)*(EIF(AA)-EIF(BB))
      GO TO 2
      CONTINUE
      XIGL=0.0
      CONTINUE
      X2IGL=(1.-EXP(-B*T))/(1.+CC*T)-XIGL/CC
      RETURN
      END

```

```

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01 JUL 74 VERSION

PACKAGE CHEM74

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C --- FUNCTION EIF (X)
C      INTEGRAL OF EXP(-X)/X FROM X TO INF
C
C      IF (X-1) > 1.2
C      EIF=-ALOG(X)+((0.00117827*X-0.00976004)*X+0.05519968)*X-0.249914
C      GO TO 3
C      EIF=EXP(-X)*(X+2.334733)*X+0.250621/((X+3.330657)*X+1.681534)/X
C      RETURN
C      END
  
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PACKAGE CHEM74

FUNCTION SHXOX (X)

```

C --- EVALUATION OF (EXP(X)-1.)/X
C
IF (ABS(X).LE.1.E-7) GO TO 1
SHXOX=(EXP(X)-1.)/X.
RETURN
CONTINUE
SHXOX=1.+X/2.
RETURN
END

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PACKAGE CHEM74

FUNCTION WR (Q+A,XU,T)

--- SOLUTION OF DX/DT=-A\*X\*\*2

IF (Q.LE.0.) GO TO 1

SQ1=SQRT(Q/A)

SQ2=SQRT(Q/A)

EXJAE=EXP(-SQ2\*\*2.\*T)

R=(XG-SQ1)/(XG\*SQ1)

WR=SQ1\*(1.+B\*EXJAE)/(1.-B\*EXQAE)

RETURN

CONTINUE

WR=X3/(1.+A\*XU\*T)

RETURN

END

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```

1702 CALL EVIL (B,F,0,Y)
1703 DO 11 LE=1,3
1704 CALL FUNG
1705 IF (NTG:EL,1) GO TO 17
1706 CALL EVIL (B,P,S,Z)
1707 DO 18 I=1,N
1708 KRK=ABS(Z(I)-Y(I))/(X(I)+1.0E-5)-1.0E-4
1709 GO TO 9
1710 CONTINUE
1711 GO TO 12
1712 DO 10 I=1,N
1713 Y(I)=Z(I)
1714 CONTINUE
1715 DT=0.5*DT
1716 IF (DT.LC.1.0E-25) STOP
1717 GO TO 7
1718 CONTINUE
1719 LL=LL+1
1720 DO 13 J=1,N
1721 U(K,J)=Z(J)
1722 X(J)=Z(J)
1723 CONTINUE
1724 CEL=CN2PL+COPL+CNOPL+CNPL+CO2PL+CUOPL+CNZAPL+CN2BPL
1725 T=1+DT
1726 U(K,30)=CEL
1727 WRITE (4,27) T, (CMD(I), I=1, NSPEC)
1728 FG=FG+V*CONV
1729 SS(K,1)=P(87)
1730 SS(K,2)=P(88)
1731 SS(K,3)=P(89)
1732 SS(K,4)=P(90)
1733 SS(K,5)=P(91)
1734 SS(K,6)=P(92)
1735 DV(K)=DT
1736 V(K)=T
1737 DT=2.0*DT
1738 IF (T-END) 14,17,17
1739 IF ((T+DT)-CMD) 16,10,15
1740 DT=END-T+1.5E-10
1741 CONTINUE
1742 CONTINUE
1743 CONTINUE
1744 CONTINUE
1745 THE FOLLOWING ARE CHEM SPECIAL PRINTS *****
1746 IF (SWTCH156..NE.1HY) GO TO 22
1747 PKINT 24
1748 DO 18 K=1,LL
1749 PRINT 25, V(K), DV(K), U(K,33), U(K,40), U(K,25), (SS(K,I), I=1, 6), TXX(K)
1750 1) CONTINUE
1751 PRINT 30
1752 DO 19 K=1,LL
1753 PRINT 34, V(K), U(K,34), U(K,35), U(K,51), U(K,22), U(K,31), U(K,38), U(K
1754 1,33), U(K,41), U(K,42), U(K,43), U(K,44)
1755 CONTINUE
1756 DO 20 K=1,LL
1757 PRINT 34, V(K), U(K,21), U(K,27), U(K,1), U(K,11), U(K,36), U(K,23), U(K,
1758 12), U(K,13), U(K,20), U(K,52), U(K,54)
1759 CONTINUE
1760 IF (CMTL:LE:0.) GO TO 22
1761 PRINT 33
1762 DO 21 K=1,LL
1763 PRINT 35, V(K), U(K,2), U(K,12), U(K,4), U(K,14), U(K,15), U(K,6)
1764 U(K,16), U(K,18), U(K,9), U(K,19)
1765 CONTINUE
1766 END OF SPECIAL PRINTS *****
1767 IF (T-END) 3,23,23
1768 CONTINUE
1769 IF (T-END) 3,23,23
1770 CONTINUE
1771 IF (T-END) 3,23,23
1772 CONTINUE
1773 IF (T-END) 3,23,23
1774 CONTINUE
1775 IF (T-END) 3,23,23
1776 CONTINUE

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1777 CONTINUE*CONV
1778 IF (INT*EQ.10) RETURN
1779 CALL RTRN (TG)
1780 IF (SM2PL*COPL+CMOPL+CM2APL+CM2BPL+CEL
1781 CMIL=CLL+CBE+CMG+CAL+CSI+CFE+CCU+CTI
1782 CMHLO=CLL+CBE+CMG+CAL+CSI+CFE+CCU+CTI
1783 PNUM=CNS+CN2+CN2A+CN2B+CN2PL+CN2APL+CN2BPL+CN2APL+CO2+CO2
1784 18+CO2SG+CO2+3.0*(CO3+CN2+CN2O+CCU2+CR2O)
1785 PRINT 36, CHG, PNUM
1786 CONTINUE
1787 RETURN
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FORMAT (57H CHEMISTRY VARIABLES ED, TIOMPR, FRION, TG, END, QQ, QXT, NFLA
19 17E13.5, 15I11) INITIAL CHARGE AND PARTICLES, 2E25.15)
FORMAT (E13.6, (6E13.6, /))
FORMAT (I3I1H1 TIME DT F-03 OS F-N02 CO F-N2
1 1 F-NO
20 FORMAT (E12.5, E10.3, 1DE11.3) N N(0) NO
FORMAT (I11, 127H N2+ N2+ NO
1 1) N2 N20 N20 U U OH
2 FORMAT (I11, 120H NO2 TIME NO2 UO+
1 2) TIME SI LI SIO LIO FE MG FEO MGO CU
2 1AL CUO) (12E11.3)
FORMAT (13E10.3) CHARGE AND PARTICLES, 2E25.15)
FORMAT (23H FINAL
END

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```
*****  
SUBROUTINE _VIL (X,P,R,Y)  
*****  
      EVIL  SOLUTION OF  $\Delta X/\Delta T = P - R * X$  FOR NUMERICAL ALGORITHM  
*****  
      DIMENSION X(1), P(1), R(1), Y(1)  
      COMMON /KEN/ CONV,DT,N  
      GO TO 3 IF N=1  
      T=DT * R(1)  
      IF (T.GT.1.5E-3) GO TO 1  
      Y(1)=X(1) * (1.-T) + DT * P(1)  
      GO TO 3  
      CONTINUE  
      RA=P(1) / R(1)  
      IF (T.LE.150.) GO TO 2  
      Y(1)=RX  
      GO TO 3  
      Y(1)=(X(1)-RA) * EXP(-T) + RA  
      CONTINUE  
      RETURN  
      END
```

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COC=
FU=FC*4LO*CU*4LO
FUOPL=FU*CU*4LO
DUOPL=DU*CU*4LO
DNZB=DALN2
DNZAPL=DALN2PL
DNZBPL=DALN2PL
IF (UU*LE*G*7) GO TO 4
CALL MDEP (TIOMPR,CON,T3)
FNS=FN5+QMS1
FND=FN0+QND1
FNP=FN0+QNP1
FNZPL=FN2PL+QNZPL
FOZPL=FN2PL+QOZPL
FNPL=FNPL+QNP1
FOPL=FOPL+QOPL
FO=FO+QFC
OO=OO+QOO
FOQ=FOQ+QOO1
FOJ=FOJ+QOO1
FOZJ=FOZJ+QOO2
FOZS=FOZS+QOO2
DNZ=DNZ+QDNZ
OOZ=OOZ+QOOZ
CONTINUE
RTURN
NTG=10
PRINT 6, T3
RETURN

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4 5 006

FORMAT (34H THIS TOO SMALL OR NEGATIVE AND = ,E14.7)  
END

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1178 222178 2 3  
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1181 222181 5 6  
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1183 222183 7 8  
1184 222184 8 9  
1185 222185 9 10  
1186 222186 10 11  
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1214 222214 38 39  
1215 222215 39 40  
1216 222216 40 41  
1217 222217 41 42  
1218 222218 42 43  
1219 222219 43 44  
1220 222220 44 45  
1221 222221 45 46  
1222 222222 46 47  
1223 222223 47 48  
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1225 222225 49 50  
1226 222226 50 51  
1227 222227 51 52  
1228 222228 52 53  
1229 222229 53 54  
1230 222230 54 55  
1231 222231 55 56  
1232 222232 56 57  
1233 222233 57 58  
1234 222234 58 59  
1235 222235 59 60  
1236 222236 60 61  
1237 222237 61 62  
1238 222238 62 63  
1239 222239 63 64  
1240 222240 64 65  
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1242 222242 66 67  
1243 222243 67 68  
1244 222244 68 69  
1245 222245 69 70  
1246 222246 70 71  
1247 222247 71 72  
1248 222248 72 73  
1249 222249 73 74  
1250 222250 74 75

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SUBROUTINE RRCT (IG)  
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PRCT  
\*\*\*\*\*  
STORES DESTRUCTION RATES OF SPECIES FOR WORRY  
CALCULATES VIBRATIONAL LEVEL FORMATION RATES OF SPECIES  
(NOT CURRENT WORRY VERSION)  
\*\*\*\*\*  
COMMON /34MDS/ HSTATE(23),NSTATE(20)  
COMMON /FU/ V(93),P(90),S(93),R(90),CHT,TGEV,QQ,TNUM,NFLAG,NTG  
COMMON /KENV/ CONV,OT,N  
COMMON /KUX/ RR(23),CMD(19),RRCT(90),NSPEC  
REAL LEN(33)  
\*\*\*\*\*  
EQUIVALENCE (CMD(1),CH), (CMD(2),CLI), (CMD(3),CBE), (CMD(4),CMG), (CMD(5),CAL), (CMD(6),CST), (CMD(7),CII), (CMD(8),CBE), (CMD(9),C  
3CU), (CMD(10),CU), (CMD(11),COH), (CMD(12),CLIO), (CMD(13),CBE), (CMD(14),CMGO), (CMD(15),CALO), (CMD(16),CSIO), (CMD(17),CTIO), (C  
3CM(18),CFEO), (CMD(19),CCUO), (CMD(20),CUOPL), (CMD(21),CNOPL), (C  
CMD(22),CMO), (CMD(23),CMO2), (CMD(24),CM2O), (CMD(25),CO), (CMD(2  
66),CCO2), (CMD(27),CO3), (CMD(28),CM2O), (CMD(30),CEL), (CMD(31),C  
7N2), (CMD(32),CN2A), (CMD(33),CN2B), (CMD(34),CNS), (CMD(35),CND),  
941), (CMD(36),CMP), (CMD(37),COP), (CMD(38),CNS), (CMD(39),CGS), (CMD(40),  
445), (CMD(41),COP), (CMD(42),COP2), (CMD(43),COP3), (CMD(44),COP4), (CMD(45),COP5), (CMD(46),COP6), (CMD(47),COP7), (CMD(48),COP8), (CMD(49),COP9), (CMD(50),COP10), (CMD(51),COP11), (CMD(52),COP12), (CMD(53),COP13), (CMD(54),COP14), (CMD(55),COP15), (CMD(56),COP16), (CMD(57),COP17), (CMD(58),COP18), (CMD(59),COP19), (CMD(60),COP20), (CMD(61),COP21), (CMD(62),COP22), (CMD(63),COP23), (CMD(64),COP24), (CMD(65),COP25), (CMD(66),COP26), (CMD(67),COP27), (CMD(68),COP28), (CMD(69),COP29), (CMD(70),COP30), (CMD(71),COP31), (CMD(72),COP32), (CMD(73),COP33), (CMD(74),COP34), (CMD(75),COP35), (CMD(76),COP36), (CMD(77),COP37), (CMD(78),COP38), (CMD(79),COP39), (CMD(80),COP40), (CMD(81),COP41), (CMD(82),COP42), (CMD(83),COP43), (CMD(84),COP44), (CMD(85),COP45), (CMD(86),COP46), (CMD(87),COP47), (CMD(88),COP48), (CMD(89),COP49), (CMD(90),COP50), (CMD(91),COP51), (CMD(92),COP52), (CMD(93),COP53), (CMD(94),COP54), (CMD(95),COP55), (CMD(96),COP56), (CMD(97),COP57), (CMD(98),COP58), (CMD(99),COP59), 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(CMD(778),COP738), (CMD(779),COP739), (CMD(780),COP740), (CMD(781),COP741), (CMD(782),COP742), (CMD(783),COP743), (CMD(784),COP744), (CMD(785),COP745), (CMD(786),COP746), (CMD(787),COP747), (CMD(788),COP748), (CMD(789),COP749), (CMD(790),COP750), (CMD(791),COP751), (CMD(792),COP752), (CMD(793),COP753), (CMD(794),COP754), (CMD(795),COP755), (CMD(796),COP756), (CMD(797),COP757), (CMD(798),COP758), (CMD(799),COP759), (CMD(800),COP760), (CMD(801),COP761), (CMD(802),COP762), (CMD(803),COP763), (CMD(804),COP764), (CMD(805),COP765), (CMD(806),COP766), (CMD(807),COP767), (CMD(808),COP768), (CMD(809),COP769), (CMD(810),COP770), (CMD(811),COP771), (CMD(812),COP772), (CMD(813),COP773), (CMD(814),COP774), (CMD(815),COP775), (CMD(816),COP776), (CMD(817),COP777), (CMD(818),COP778), (CMD(819),COP779), (CMD(820),COP780), (CMD(821),COP781), (CMD(822),COP782), (CMD(823),COP783), (CMD(824),COP784), (CMD(825),COP785), (CMD(826),COP786), (CMD(827),COP787), (CMD(828),COP788), (CMD(829),COP789), 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PACKAGE CHEM74

RETURN  
END

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N 225-  
N 226-











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RETURN

FORMAT (E13.6,(6E13.6,/))  
FORMAT (1A,6K,TIME=,I5,3H KATIME=,I5)  
END

```

2685 SUBROUTINE GRID (NPROB)
2686 COMMON /PLOT/IXN,IPX
2687 COMMON /TITLE/ITL(7)
2688 DIMENSION SPEC(5)
2689 DATA (SPEC(I),I=1,5)/5H H,5H LI,5H 3E,5H MG,5H AL,5H MG
2690 10,5H ALO,5H N2,5H CU,5H FEU,5H CUO,5H UOPL,5H NO,5H NO,5H
2691 35H N2,5H N2O,5H CO,5H CO2,5H H2O,5H NP,5H OP,5H
2692 45 OD,5H N2A,5H N2B,5H NS,5H NP,5H NP,5H NP,5H
2693 5 OSM,5H N2PL,5H O2PL,5H NPL,5H OPL,5H MN2,5H PL,5H
2694 6 NP=NPPL(NPROB)
2695 IX=XSKL*8.
2696 CALL NUMBER(1,16,3)
2697 CALL PLOT(IP)
2698 TIB=8.0/IX
2699 DO 2 I=1,IX+1.
2700 1 LK=1.0
2701 PJ=ALOG10(LK+1.)
2702 CALL PLOT(IP)*TIB
2703 CALL PLOT(IP)
2704 CALL PLOT(IP)
2705 CALL NUMBER(IX+1,1,3)
2706 CALL PLOT(IP)
2707 CALL PLOT(IP)
2708 CALL PLOT(IP)
2709 CALL NUMBER(IX+1,1,3)
2710 CALL PLOT(IP)
2711 CALL PLOT(IP)
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